



EXPERIMENT 1: DETERMINING MOLAR VOLUME

Data Sheet

Table 3: Amount of H_2O Displaced for Each Trial-Grams of Solute Used

Trial	Barometric Pressure (atm)	Temperature ($^{\circ}\text{C}$)	Mass of Chemical (g)	Water Displaced (mL)
Trial 1 NaHCO_3	.99	19.6	1.0	83
Trial 2 NaHCO_3	.99	20.2	.9	68
Trial 3 Na_2CO_3	.99	20.0	1.3	96
Trial 4 Na_2CO_3	.99	20.0	1.0	105

Post-Lab Questions

1. What is the average experimental volume per mole of carbon dioxide calculated for sodium carbonate and sodium bicarbonate?
2. List two sources of error in your experiment. What did they affect? Explain how they could have been fixed.
3. Identify two assumptions that must be true when using the ideal gas equation that are not true for real gases.
4. Why is the air in a hot air balloon heated?

Determining The Molar Volume Of A Gas Lab Worksheet

Peipei Pang



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ChemDiscovery Teacher Edition Olga I. Agapova,2002 Merrill Chemistry-Lab.Manual Smoot,1994-07

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