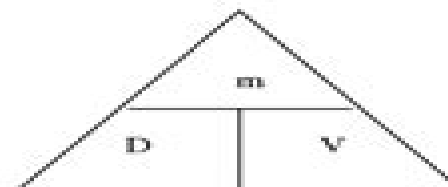


SNC 1D0
Unit: Chemistry
Density Calculations Worksheet

Name: _____
 Date: _____

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

UNITS OF DENSITY
 g/cm^3 or g/mL



1. Find the unknown quantity.

a. $D = 3 \text{ g/mL}$
 $V = 100 \text{ mL}$
 $m = ?$

300 g

b. $D = ?$
 $V = 950 \text{ mL}$
 $m = 95 \text{ g}$

0.10 g/mL

c. $D = 0.5 \text{ g/cm}^3$
 $V = ?$
 $m = 20 \text{ g}$

40 cm^3

2. Find the unknown quantity (CONVERT FIRST to g or mL).

a. $D = 24 \text{ g/mL}$
 $V = 1.2 \text{ L} = \underline{\hspace{1cm}} 1200 \underline{\hspace{1cm}} \text{ mL}$
 $M = ?$

28800 g = 28.8 kg

b. $D = ?$
 $V = 100 \text{ mL}$
 $M = 1.5 \text{ kg} = \underline{\hspace{1cm}} 1500 \underline{\hspace{1cm}} \text{ g}$

15 g/mL

c. $D = ?$
 $V = 0.52 \text{ L} = \underline{\hspace{1cm}} 520 \underline{\hspace{1cm}} \text{ mL}$
 $M = 500 \text{ mg} = \underline{\hspace{1cm}} 0.5 \underline{\hspace{1cm}} \text{ g}$

0.00096 g/mL

WORD PROBLEMS

1. A block of aluminum occupies a volume of 15.0 mL and has a mass of 40.5 g. What is its density?

2.70 g/mL

2. Mercury metal is poured into a graduated cylinder that holds exactly 22.5 mL. The mercury is used to fill the cylinder has a mass of 306.0 g. From this information, calculate the density of mercury.

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M Woodhall



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