

Experiment 7
The Solution is Dilution
REPORT

1. Fill in this table using the instructions in the lab manual and from video:

Test Tube #	mL Stock Solution	mL DI water	Total Volume of solution (mL)	Molarity of solution, M
1	10.0	0.0		0.20
3	8.0	2.0		0.16
3	6.0	4.0		0.12
4	4.0	6.0		0.08
5	2.0	8.0		0.04

2. Complete the "Student analysis of colorimeter data and unknown" page of this experiment.

What is the concentration of the unknown solution of nickel(II) chloride? 0.143

Absorbance in video = 0.425

a = εcl

Questions needing solutions:

1. Does the absorbance of a solution depend on the concentration? yes
(Explain your reasoning.)

2. Use the lab manual to find the wavelength of light absorbed and the color wheel (previous experiment) to fill in the table below.

Solution	Color of light absorbed	Color of solution
$\text{Co}(\text{NO}_3)_2(\text{aq})$		
$\text{NiCl}_2(\text{aq})$		

What is the relationship between the color of light absorbed by the solution and the color of the solution? (Explain below)

3. Suppose 25.00 g of $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ (sucrose) is added to a 50.00 mL volumetric flask. DI water is added to the flask until the total volume of the solution is 50.00 mL. What is the molarity of the resulting solution? answer: _____
4. If 25.00 mL of the resulting solution is diluted with water to a total volume of 150 mL. What is the concentration of the diluted solution? answer: _____
(Show calculation here)

Dilution Of Solution Lab Report

SJ Ball



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