

Physical Science- Chapter 24 Test Study Guide KEY

1. _____ - the substances that are about to react.
Reactants
2. _____ - the new substances produced.
Products
3. Law of Conservation of Mass- the French alchemist Antoine Lavoisier established that the mass of the products always _____ the mass of the reactants.
Equals
4. _____ - a way to describe a chemical reaction using chemical formulas and other symbols.
Chemical Equation
5. (s) means? _____
Solid
6. (aq) means? _____
Aqueous, dissolved in water
7.  means? _____
Yields, produces, forms
8. In a chemical reaction atoms can be rearranged but never _____ or _____.
Created, destroyed
9. The numbers to the left of formulas for reactants and products are called _____.
Coefficients
10. _____ - represents the number of units of each substance taking part in a reaction.
Coefficients
11. The following underlined numbers are called _____.
$$\text{NiCl}_2 (\text{aq}) + \underline{2}\text{NaOH} (\text{aq}) \rightarrow \text{Ni}(\text{OH})_2 (\text{s}) + \underline{2}\text{NaCl} (\text{aq})$$

Coefficients
12. Iron exposed to _____ and _____ corrodes or rusts forming iron (III) oxide, rust can seriously damage iron structures because it crumbles and exposes more iron to the air.
Air, moisture
13. Copper corrodes when exposed to the air forming a blue-green coating called a _____.
Patina
14. _____ - has the same number of atoms of each element on both sides of the equation.
Balanced chemical equation

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Table of Contents Ch 24 Study Guide For Content Mastery

1. Understanding the eBook Ch 24 Study Guide For Content Mastery
 - The Rise of Digital Reading Ch 24 Study Guide For Content Mastery
 - Advantages of eBooks Over Traditional Books
2. Identifying Ch 24 Study Guide For Content Mastery
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ch 24 Study Guide For Content Mastery
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ch 24 Study Guide For Content Mastery
 - Personalized Recommendations
 - Ch 24 Study Guide For Content Mastery User Reviews and Ratings
 - Ch 24 Study Guide For Content Mastery and Bestseller Lists
5. Accessing Ch 24 Study Guide For Content Mastery Free and Paid eBooks

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- Ch 24 Study Guide For Content Mastery eBook Subscription Services
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 - ePub, PDF, MOBI, and More
 - Ch 24 Study Guide For Content Mastery Compatibility with Devices
 - Ch 24 Study Guide For Content Mastery Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ch 24 Study Guide For Content Mastery
 - Highlighting and Note-Taking Ch 24 Study Guide For Content Mastery
 - Interactive Elements Ch 24 Study Guide For Content Mastery
- 8. Staying Engaged with Ch 24 Study Guide For Content Mastery
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ch 24 Study Guide For Content Mastery
- 9. Balancing eBooks and Physical Books Ch 24 Study Guide For Content Mastery
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ch 24 Study Guide For Content Mastery
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ch 24 Study Guide For Content Mastery
 - Setting Reading Goals Ch 24 Study Guide For Content Mastery
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch 24 Study Guide For Content Mastery
 - Fact-Checking eBook Content of Ch 24 Study Guide For Content Mastery
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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