

Chapter 2 Study Guide

CHAPTER 2 The Chemical Basis of Life I: Atoms, Molecules, and Water

- Describe the location, charge and mass of protons, neutrons and electrons, within an atom.**
 - Neutrons and Protons are densely packed in the Nucleus.
 - Electrons move about the Nucleus in Orbitals.
 - Neutrons and Protons have a mass of one.
- Be able to calculate the Atomic Weight of an atom/element, if the number of protons, electrons and neutrons is known.**
 - Atomic Mass Number = Protons + Neutrons
- Be able to calculate the Atomic Number of an atom/element, if the number of protons, electrons and neutrons is known.**
 - Atomic Number = Number of Protons
- Be able to write the symbols, atomic numbers, and mass numbers of the biologically important elements (C, H, O, & N).**
 - C → Carbon, 12 Protons, 12.011 Mass
 - H → Hydrogen, 1 Proton, 1.008 Mass
 - O → Oxygen, 8 Protons, 15.999 Mass
 - N → Nitrogen, 7 Protons, 14.007 Mass
- Be able to explain how isotopes of an element differ in terms of the number of electrons, neutrons and protons.**
 - Isotopes differ in the number of neutrons in the Nucleus.
 - The number of Protons does not change.
 - The number of Electrons does not change.
- Be able to correctly place any number of electrons in an atom's first two energy level shells.**
 - First Energy Level requires 2 Electrons.
 - Any Energy Levels after the first require 8 Electrons.
- Differentiate between atoms, molecules and ions.**
 - Atoms are single neutral particles.
 - Molecules are neutral particles made of two or more atoms bonded together.
 - Ions are positively or negatively charged particles (atoms)
- Differentiate between molecules and compounds.**
 - All compounds are molecules, but not all molecules are compounds.
 - Molecules are made up of any two or more atoms bonded together, but Compounds are made up of two or more different atoms bonded together.
- What is a valence electron?**
 - Valence Electrons are Electrons located in the outermost Energy Shell of an atom. Since they are the farthest from the Nucleus, they can participate in bonds and reactions, therefore Valence Electrons determine an atom's reactivity, electronegativity, and the number of bonds it can form.
- Describe a covalent chemical bond, and what holds the two atoms together.**
 - A Covalent Bond is formed through the equal sharing of valence Electrons in two atoms, with the sharing process enabling each atom to fill its outer valence electrons.
- Differentiate between a polar and non-polar covalent bond.**
 - A Polar Covalent Bond results if the Electronegativity difference from a bond is between 0.5-1.6, meaning one end of the bond is partly negative while the other is partly positive.
 - A Non-Polar Covalent Bond results if the Electronegativity difference from a bond is between 0.2-0.5.
- What is a Hydrogen Bond, and what holds the two atoms together?**
 - A Hydrogen Bond occurs only with Polar Molecules, as the partial opposite charges of the molecules will attract one another, usually placing a Hydrogen atom as the bridge between the molecules.
- What is an ionic bond, and what holds the two atoms together.**
 - An Ionic Bond is a bond that is formed from the transfer of Electrons from one atom to another, therefore making one element have more Electrons than Protons and the other have fewer Electrons than Protons.
 - Essentially, the more Electronegative atom strips Electrons from another atom to complete its outer-most valence shell.
- What is a van der Waals interaction (bond) and what holds the two atoms together?**

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Study Guide and Full Solutions Manual Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all end of chapter problems in the textbook

Study Guide and Partial Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry Susan McMurry, 1996 This internationally acclaimed detective series is just the thing for lovers of those Number One Ladies looking for a darker more realistic view of Botswana Sue Baker Publishing News

Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

Riding on a Ray of Light Krishnagopal Dharani, 2020-04-07 A Conceptual Breakthrough in Our Understanding of Fundamental Nature of Matter and Energy A lot of questions have bothered science for a long time What is a photon Why does light behave both like a particle and a wave How does light transform into matter What is gravity What is Big Bang and what came before it The list is endless Riding on a Ray of Light describes a working model called the Negentropic Model which describes the fundamental

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Study Guide to Chemistry and Life John William Hill,Dorothy M. Feigl,1978 *Foundations of College Chemistry* Morris Hein,Susan Arena,Cary Willard,2023 Foundations of College Chemistry 16th edition presents chemistry as a modern vital subject and is designed to make introductory chemistry accessible to all beginning students It is intended for students who have never taken a chemistry course or those who had a significant interruption in their studies but plan to continue with the general chemistry sequence The central focus is to make chemistry interesting and understandable and teach students the problem solving skills they will need This International Adaptation offers new and updated content with improved presentation of all course material It builds on the strengths of previous editions including clear explanations and step by step problem solving The material emphasizes real world applications of chemistry as the authors develop the principles that form the foundation for the further study of chemistry There is new and expanded coverage of polarizing power and polarizability Fajans rules collision number and mean free path abnormal molecular masses and van t Hoff factor and

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Introduction to General, Organic, and Biochemistry Study Guide William Scovell,2006-03 By William M Scovell This resource helps students organize their study time and guides them through the topics in a systematic way Each chapter of the text is covered by an introduction a list of review topics section by section study suggestions and questions a list of key terms and a practice exam with worked out answers **Study Guide to Accompany Introduction to General, Organic, & Biochemistry** Frederick A. Bettelheim,William M. Scovell,1998 **Hulme branch lending library. Class list of works on natural science, useful arts, and fine arts** Manchester publ. libr,1905

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