

Chapter 10: Chemical Quantities

Section 1 Notes – The Mole: A Measurement of Matter (Read pages 287-296)

Mole

- chemist _____ for number of _____
- equals _____ number = _____

Representative Particles

- atoms – _____
- molecules – _____
- formula units – _____
- ions – _____

Conversions

- _____ mole equals
 - _____ atoms, molecules, formula units, ions, things

Practice Problems (Do Together)

3. How many moles is 2.80×10^{24} atoms of Si?
4. How many moles if 2.17×10^{23} representative particles of bromine?
5. How many atoms are in 1.14 mol SO_2 ?
6. How many moles are in 4.65×10^{24} molecules of NO_2 ?

Additional Practice Problems

1. How many moles are equal to 9.03×10^{24} atoms of mercury?
2. How many atoms are in 1.00 mole of sacrose, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$?
3. How many atoms of O are in 3.65 moles of $\text{C}_{12}\text{H}_{22}\text{O}_{11}$?

Chapter 10 Chemical Quantities Guided Practice Problems Answers

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Armstrong**



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