

# 11 Chemical Reactions

## Big idea REACTIONS

### 11.1 Describing Chemical Reactions

For students using the Foundation edition, assign problems 1–17.

**Essential Understanding** Chemical reactions are represented by balanced chemical equations.

#### Lesson Summary

**Introduction to Chemical Equations** A chemical equation uses symbols, and sometimes words, to show the reactants and products of a chemical reaction.

- ▶ A skeleton equation uses chemical formulas to represent reactants and products, but it does not indicate the relative amounts of each.
- ▶ Anything that enters into a reaction, such as heat or a catalyst, but is not a reactant or product is shown above or below the yields arrow in the equation.

**Balancing Chemical Equations** Chemical equations are balanced to show that mass is conserved during chemical reactions.

- ▶ A balanced equation shows the relative amounts of reactants and products, and it contains equal numbers of each type of atom on both sides of the equation.
- ▶ Chemical equations are balanced by using coefficients in front of the chemical formulas for the reactants and the products in a skeleton equation.



#### BUILD Math Skills

**Balancing Equations** All chemical equations must be balanced because of the law of conservation of mass, which states that matter cannot be created or destroyed. So, the number of atoms that you start with at the beginning of the reaction must equal the number of atoms that you end up with.

For example, the reaction  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$  follows the law of conservation of mass because you start with 2 magnesium atoms and 2 oxygen atoms and you end up with 2 magnesium atoms and 2 oxygen atoms. You can think of the  $\rightarrow$  as an  $=$  sign.



Turn the page to learn more about balancing equations.

# Chapter 11 Section 11 1 Chemical Reactions Answer Key

**Lei Huang**



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*Organic Chemistry, Fourth Edition* K. Peter C. Vollhardt, Neil E. Schore, 2003 New edition of the acclaimed organic chemistry text that brings exceptional clarity and coherence to the course by focusing on the relationship between structure and function

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*Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

*Foundations of Chemistry* Eleanor Dantzler Siebert, 1982

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**Modules** McDougal Littell Incorporated, 2005

Chemistry Oxtoby, 1994

**The Northwestern Druggist**, 1905

*Student Study Guide to Accompany Petrucci's General Chemistry, 3rd. Ed* Robert K. Wismer, 1982

*Introduction to Chemical Reaction Engineering and Kinetics* Ronald W. Missen, Charles A. Mims, Bradley A. Saville, 1999 Solving problems in chemical reaction engineering and kinetics is now easier than ever As students read through this text they ll find a comprehensive introductory treatment of reactors for single phase and multiphase systems that exposes them to a broad range of reactors and key design features They ll gain valuable insight on reaction kinetics in relation to chemical reactor design They will also utilize a special software package that helps them quickly solve systems of algebraic and differential equations and perform parameter estimation which gives them more time for analysis Key Features Thorough coverage is provided on the relevant principles of kinetics in order to develop better designs of chemical reactors E Z Solve software on CD ROM is included with the text By utilizing this software students can have more time to focus on the development of design models and on the interpretation of calculated results The software also facilitates exploration and discussion of realistic industrial design problems More than 500 worked examples and end of chapter problems are included to help students learn how to apply the theory to solve design problems A web site [www.wiley.com/college/missen](http://www.wiley.com/college/missen) provides additional resources including sample files demonstrations and a description of the E Z Solve software

*The Software Encyclopedia*, 1988

**The Pharmaceutical Journal and Transactions**, 1878

Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

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