

## Study Card: Enthalpy Change

### What is Enthalpy?

- Enthalpy is the total energy possessed by a chemical system
- Represented as  $H$  and has units of joules (J) or kilojoules (kJ)
- Includes potential energy associated with particle interactions,
- Includes kinetic energy associated with particle movement
- Includes energy required to make room for the system in the surroundings (often referred to as  $P \cdot V$ )
- A chemical reaction causes an **enthalpy change** ( $\Delta H$ )
- At a constant pressure, the  $\Delta H$  equals the heat ( $Q$ ) exchanged between the system and the surroundings

### Enthalpy Level Diagrams

Enthalpy changes in reactions are often represented by enthalpy level diagrams. The enthalpy of the reaction system is displayed as a function of some indicator of the reaction's progress.

#### Exothermic Reactions

$\Delta H$  is negative ( $<0$ )

$H_{\text{reactants}} > H_{\text{products}}$

Heat released by system

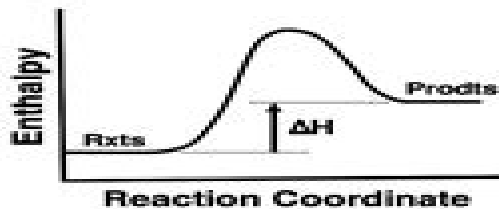


#### Endothermic Reactions

$\Delta H$  is positive ( $>0$ )

$H_{\text{reactants}} < H_{\text{products}}$

Heat absorbed by system



### Thermochemical Equations

Every chemical reaction has a unique enthalpy change or  $\Delta H$  value. Enthalpy change is a **stoichiometric quantity**; it depends on how much reactants are involved in the reaction. It is often included in the chemical equation for the reaction:



(Translation: 890 kJ of energy are produced by the reaction.  $\Delta H = -890 \text{ kJ}$ )



(Translation: 572 kJ of energy are required by the reaction.  $\Delta H = +572 \text{ kJ}$ )

If the coefficients are changes, the heat term changes accordingly. If the reactants and products are reversed, the heat term changes side accordingly:



### System Diagrams

#### Exothermic Reaction

$\Delta H < 0$

$H_{\text{rxts}} > H_{\text{prodts}}$



Heat released by system.

#### Endothermic Reaction

$\Delta H > 0$

$H_{\text{rxts}} < H_{\text{prodts}}$



Heat absorbed by system.

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[www.physicsclassroom.com](http://www.physicsclassroom.com)

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# Causes Of Change Enthalpy Concept Review Answer

**SILBERBERG**



## Causes Of Change Enthalpy Concept Review Answer:

*Chemistry3* Andrew Burrows, John Holman, Simon Lancaster, Andrew Parsons, Tina Overton, Gwen Pilling, Gareth Price, 2021 Chemistry is widely considered to be the central science it encompasses concepts on which all other branches of science are developed Yet for many students entering university gaining a firm grounding in chemistry is a real challenge *Chemistry3* responds to this challenge providing students with a full understanding of the fundamental principles of chemistry on which to build later studies Uniquely amongst the introductory chemistry texts currently available *Chemistry3*'s author team brings together experts in each of organic inorganic and physical chemistry with specialists in chemistry education to provide balanced coverage of the fundamentals of chemistry in a way that students both enjoy and understand The result is a text that builds on what students know already from school and tackles their misunderstandings and misconceptions thereby providing a seamless transition from school to undergraduate study Written with unrivalled clarity students are encouraged to engage with the text and appreciate the central role that chemistry plays in our lives through the unique use of real world context and photographs *Chemistry3* tackles head on two issues pervading chemistry education students mathematical skills and their ability to see the subject as a single unified discipline Instead of avoiding the maths *Chemistry3* provides structured support in the form of careful explanations reminders of key mathematical concepts step by step calculations in worked examples and a Maths Toolkit to help students get to grips with the essential mathematical element of chemistry Frequent cross references highlight the connections between each strand of chemistry and explain the relationship between the topics so students can develop an understanding of the subject as a whole Digital formats and resources *Chemistry3* is available for students and institutions to purchase in a variety of formats and is supported by online resources The e book offers a mobile experience and convenient access along with functionality tools navigation features and links that offer extra learning support [www.oxfordtextbooks.co.uk/ebooks](http://www.oxfordtextbooks.co.uk/ebooks) The e book also features interactive animations of molecular structures screencasts in which authors talk step by step through selected examples and key reaction mechanisms and self assessment activities for each chapter The accompanying online resources will also include for students DT Chapter 1 as an open access PDF DT Chapter summaries and key equations to download to support revision DT Worked solutions to the questions in the book The following online resources are also provided for lecturers DT Test bank of ready made assessments for each chapter with which to test your students DT Problem solving workshop activities for each chapter for you to use in class DT Case studies showing how instructors are successfully using *Chemistry3* in digital learning environments and to support innovative teaching practices DT Figures and tables from the book      **Holt Chemistry** Holt Rinehart & Winston, 2003-01-24

**Chemistry<sup>3</sup>** Andrew Burrows, Andrew Parsons, Gwen Pilling, Gareth Price, 2013-03-21 New to this Edition      Holt Chemistry Ralph Thomas Myers, 2004      *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

*Campbell Biology Australian and New Zealand Edition* Jane B. Reece, Noel Meyers, Lisa A. Urry, Michael L. Cain, Steven A.

Wasserman, Peter V. Minorsky, 2015-05-20 Over nine successful editions CAMPBELL BIOLOGY has been recognised as the world's leading introductory biology textbook The Australian edition of CAMPBELL BIOLOGY continues to engage students with its dynamic coverage of the essential elements of this critical discipline It is the only biology text and media product that helps students to make connections across different core topics in biology between text and visuals between global and Australian New Zealand biology and from scientific study to the real world The Tenth Edition of Australian CAMPBELL BIOLOGY helps launch students to success in biology through its clear and engaging narrative superior pedagogy and innovative use of art and photos to promote student learning It continues to engage students with its dynamic coverage of the essential elements of this critical discipline This Tenth Edition with an increased focus on evolution ensures students receive the most up to date accurate and relevant information *Thermodynamics with Chemical Engineering Applications*

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*Introduction and Computations for Gases* American Institute of Chemical Engineers, 1981 **Applied Mechanics Reviews**, 1977 Annual Review of Biophysics and Bioengineering L. J. Mullins, 1983 **Russian Chemical Reviews**, 1995 Mathematical Reviews, 2005 Chemical Society Reviews Chemical Society (Great Britain), 1972 Publishes short introductory overviews of topics of current interest across the chemical sciences The philosophy of the journal is to introduce chemists not familiar with the field under discussion to the most recent thinking and developments Authors are asked to provide background and historical context to the review as well as a clear but concise guide to current thought **Reviews in Partial Differential Equations, 1980-86, as Printed in Mathematical Reviews**, 1988 Thermodynamic Properties of Aqueous Solutions Organic Substances V. P. Belousov, Mikhail I. U. r'evich Panov, 1994 Thermodynamic Properties of Aqueous Solutions of Organic Substances discusses the structure of aqueous solutions of organic substances and the intermolecular reactions in them presenting experimental data modern concepts concerning the properties of these solutions and the results of computer simulation The book offers an in depth study of the properties of maximally dilute aqueous solutions of polar and nonpolar organic molecules as well as the specific enthalpies of mixing The Addendum contains

experimental data on the thermodynamic properties of infinitely dilute solutions      CHEMISTRY SILBERBERG,2003

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William R. Robinson,Jerome D. Odom,Henry Fuller Holtzclaw,1992

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