

# Mole Conversions Practice Answer Key

Molar Mass – Calculate the Molar Mass for the following.

- What is the molar mass of Magnesium metal?

$$24.305 \text{ g/mol}$$

- Calculate the molar mass of NaCl

$$\begin{aligned} \text{Na} &: 22.990 \times 1 = 22.990 \\ \text{Cl} &: 35.453 \times 1 = 35.453 \end{aligned} \Rightarrow 58.443 \text{ g/mol}$$

- What is the molar mass of NaOH?

$$\begin{aligned} \text{Na} &: 1 \times 22.990 = 22.990 \\ \text{O} &: 1 \times 15.999 = 15.999 \\ \text{H} &: 1 \times 1.008 = 1.008 \end{aligned} \Rightarrow 39.997 \text{ g/mol}$$

Conversions – Calculate the following the conversions. Use your Mole Map to help you!

- Determine the number of atoms in 0.50 moles He.

$$0.5 \text{ mol He} \times \frac{6.02 \times 10^{23} \text{ atoms He}}{1 \text{ mol He}} = 3.01 \times 10^{23} \text{ atoms He}$$

- Determine the number of moles in 80.00 g O<sub>2</sub>

$$0.2 \times 15.999 = 31.998 \text{ g/mol}$$

$$\Rightarrow 1 \text{ mol} = 31.998 \text{ g}$$

- Determine the volume of 0.24 mol of NO<sub>2</sub>.

$$0.24 \text{ mol NO}_2 \times \frac{22.4 \text{ L NO}_2}{1 \text{ mol NO}_2} = 5.376 \text{ L NO}_2$$

$$80.00 \text{ g O}_2 \times \frac{1 \text{ mol O}_2}{31.998 \text{ g O}_2} = 25.00 \text{ g O}_2$$

- Determine the number of moles in  $7.5 \times 10^{19}$  formula units of KNO<sub>3</sub>

$$7.5 \times 10^{19} \text{ f.u. KNO}_3 \times \frac{1 \text{ mol KNO}_3}{6.02 \times 10^{23} \text{ f.u. KNO}_3} = 1.2458 \times 10^{-4} \text{ f.u. KNO}_3$$

- Determine the number of moles in 202.22g KNO<sub>3</sub>

$$\begin{aligned} \text{K} &: 1 \times 39.098 = 39.098 \\ \text{N} &: 1 \times 14.007 = 14.007 \\ \text{O} &: 3 \times 15.999 = 47.997 \end{aligned} \Rightarrow 101.102 \text{ g/mol} \Rightarrow 1 \text{ mol} = 101.102 \text{ g}$$

- Determine the number of moles in 32.5L of N<sub>2</sub> at STP

$$32.5 \text{ L N}_2 \times \frac{1 \text{ mol N}_2}{22.4 \text{ L N}_2} = 1.451 \text{ mol N}_2$$

$$202.22 \text{ g KNO}_3 \times \frac{1 \text{ mol KNO}_3}{101.102 \text{ g KNO}_3} = 2.00 \text{ mol KNO}_3$$

- Determine the number of moles in  $3.4 \times 10^{26}$  formula units of NaCl.

$$3.4 \times 10^{26} \text{ f.u. NaCl} \times \frac{1 \text{ mol NaCl}}{6.02 \times 10^{23} \text{ f.u. NaCl}} = 569.784 \text{ f.u. NaCl}$$

- Determine the moles in 36.5L of carbon monoxide

$$36.5 \text{ L CO} \times \frac{1 \text{ mol CO}}{22.4 \text{ L CO}} = 1.629 \text{ mol CO}$$

# Challenging Mole Conversion Problems Key Answers With Steps

**Norbert T. Porile**



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The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/wink/POC/PowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/wink/POC/PowerPoint_Final.ppt) 832KB

**Essentials of Chemical Reaction Engineering** H. Scott Fogler, 2010-11-02 Learn Chemical Reaction Engineering through Reasoning Not Memorization Essentials of Chemical Reaction Engineering is a complete yet concise modern introduction to chemical reaction engineering for undergraduate students While the classic Elements of Chemical Reaction Engineering Fourth Edition is still available H Scott Fogler distilled that larger text into this volume of essential topics for undergraduate students Fogler's unique way of presenting the material helps students gain a deep intuitive understanding of the field's essentials through reasoning not memorization He especially focuses on important new energy and safety issues ranging from solar and biomass applications to the avoidance of runaway reactions Thoroughly classroom tested this text reflects feedback from hundreds of students at the University of Michigan and other leading universities It also provides new resources to help students discover how reactors behave in diverse situations Coverage includes Crucial safety topics including ammonium nitrate CSTR explosions nitroaniline and T2 Laboratories batch reactor runaways and SACHE CCPS resources Greater emphasis on safety following the recommendations of the Chemical Safety Board CSB 2 case studies from plant explosions and two homework problems which discuss another explosion Solar energy conversions chemical thermal and catalytic water spilling Algae production for biomass Mole balances batch continuous flow and industrial reactors Conversion and reactor sizing design equations reactors in series and more Rate laws and stoichiometry Isothermal reactor design conversion and molar flow rates Collection and analysis of rate data

Multiple reactions parallel series and complex reactions membrane reactors and more Reaction mechanisms pathways bioreactions and bioreactors Catalysis and catalytic reactors Nonisothermal reactor design steady state energy balance and adiabatic PFR applications Steady state nonisothermal reactor design flow reactors with heat exchange **Environmental Challenges and Greenhouse Gas Control for Fossil Fuel Utilization in the 21st Century** M. Mercedes Maroto-Valer, Chunshan Song, Yee Soong, 2012-12-06 As we are moving ahead into the 21st century our hunger for cost effective and environmentally friendly energy continues to grow The Energy Information Administration of US has forecasted that only in the first two decades of the 21st century our energy demand will increase by 60% compared to the levels at the end of the 20th century Fossil fuels have been traditionally the major primary energy sources worldwide and their role is expected to continue growing for the forecasted period due to their inherent cost competitiveness compared to non fossil fuel energy sources However the current fossil energy scenario is undergoing significant transformations especially to accommodate increasingly stringent environmental challenges of contaminants like sulfur dioxide nitrogen oxides or mercury while still providing affordable energy Furthermore traditional fossil fuel utilization is inherently plagued with greenhouse gas emissions from combustion especially carbon dioxide from stationary sources as well as from mobile sources Should worldwide government policies dictate a reduction of greenhouse gas emissions such as proposed by the Kyoto Protocol and the implementation of carbon taxes fossil fuels would lose their significant competitive appeal in favor of nuclear energy and renewable energy sources However the current non fossil fuel energy share of the worldwide energy market is merely below 15% and therefore it is more likely that fossil fuel energy producers would adapt to the new requirements by developing and implementing emission control technologies and emission trades among other strategies *The Software Encyclopedia* 2001, 2001 [Biomimetic Approaches in Engineering Practice](#) Stanislav Kolisnychenko, 2018-06-18 Aggregated Book

**Evaluation in the Post-Truth World** Mita Marra, Karol Olejniczak, Arne Paulson, 2024-04-16 Evaluation in the Post Truth World explores the relationship between the nature of evaluative knowledge the increasing demand in decision making for evaluation and other forms of research evidence and the post truth phenomena of antiscience sentiments combined with illiberal tendencies of the present day Rather than offer a checklist on how to deal with post truth the experts found herein wish to raise awareness and reflection throughout policy circles on the factors that influence our assessment and policy related work in such a challenging environment Journeying alongside the editor and contributors readers benefit from three guiding questions to help identify specific challenges but tools to deal with such challenges How are policy problems conceptualized in the current political climate What is the relationship between expertise and decision making in today's political circumstances How complex has evaluation become as a social practice Evaluation in the Post Truth World will benefit evaluation practitioners at the program and project levels as well as policy analysts and scholars interested in applications of evaluation in the public policy domain Chapters 6 and 11 of this book are freely available as a downloadable

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Modern University Chemistry, Custom Pub Norbert T. Porile, 1993

**Chemical Engineering Education**, 1994

**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998

*Current Engineering Practice*, 1986

**Academic Computing**, 1987

**Foundations of College Chemistry** Morris Hein, Susan Arena, Cary

Willard, 2016-08-02 This text is an unbound three hole punched version Used by over 750 000 students Foundations of College Chemistry Binder Ready Version 15th Edition is praised for its accuracy clear no nonsense approach and direct writing style Foundations direct and straightforward explanations focus on problem solving making it the most dependable text on the market Its comprehensive scope proven track record outstanding in text examples and problem sets were all designed to provide instructors with a solid text while not overwhelming students in a difficult course Foundations fits into the prep intro chemistry courses which often include a wide mix of students from science majors not yet ready for general chemistry allied health students in their 1st semester of a GOB sequence science education students for elementary school teachers to the occasional liberal arts student fulfilling a science requirement Foundations was specifically designed to meet this wide array of needs

**Solving Problems in Chemistry** Rod O'Connor, Charles Mickey, Alton Hassell, 1977

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