

Gases Study Guide**Test Format**

Multiple Choice
Gas Law Problems
Gas Stoich Problems
Short Answer

Equations & Conversions

$$P = \frac{F}{A}$$

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$

$$PV = nRT$$

$$R = 0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$1 \text{ atm} = 760 \text{ mmHg} = 101.3 \text{ kPa}$$

$$\frac{\text{Rate of Effusion of L}}{\text{Rate of Effusion of H}} = \sqrt{\frac{\text{Molar Mass of H}}{\text{Molar Mass of L}}}$$

$$K = ^\circ\text{C} + 273 \quad 0.001 \text{ kL} = 1 \text{ L} = 1,000 \text{ mL}$$

$$\text{At STP, 1 mole} = 22.4 \text{ L} \quad \text{KHDSdcm}$$

1. A box weighs 180 lbs and has a base that is 6 inches by 9 inches. What is the pressure this box exerts on the surface it is sitting on?

$$P = \frac{F}{A} \quad \frac{180 \text{ lbs}}{54 \text{ in}^2} = \boxed{3.3 \text{ psi}}$$

$$54 \text{ in}^2 = \text{area}$$

2. What is atmospheric pressure?

The pressure exerted by the collisions of gases in the air surrounding Earth (N_2 , O_2 , etc).

3. Convert 1.088 atm to:

a. mm Hg $\frac{1.088 \text{ atm}}{1} \cdot \frac{760 \text{ mmHg}}{1 \text{ atm}} = \boxed{826.9 \text{ mmHg}}$

b. kPa $\frac{1.088 \text{ atm}}{1} \cdot \frac{101.3 \text{ kPa}}{1 \text{ atm}} = \boxed{110.2 \text{ kPa}}$

4. What is a barometer and who invented it?

A tool that measures room/air pressure. Invented by Torricelli.



5. How are temperature and kinetic energy related?

Direct. As temp \uparrow , KE \uparrow , molecules move faster.
As temp \downarrow , KE \downarrow , molecules move slower.

6. Briefly describe the five assumptions of the kinetic-molecular theory.

Gases have a lot of tiny particles that are always in motion that is random, fast, and continuous. Particles are far apart without any forces of attraction, so their collisions are elastic \rightarrow bouncing off of each other/walls. The speed of the molecules change with temperature.

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Paul B. Kelter



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