

Chapter 10 Practice Problems & Test Review

Chemical Quantities **Answer Sheet**

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1. How many moles is 750 billion atoms H in water? (1 billion = 1×10^9)

$$\text{H} = 1.25 \times 10^{-12} \text{ H moles} \quad \text{H}_2\text{O} = 6.25 \times 10^{-13} \text{ moles H}_2\text{O}$$

2. How many carbon atoms are in a 9 mole gasoline? (gasoline = C_8H_{18})

$$4.32 \times 10^{25} \text{ C atoms}$$

3. What is the molar mass of Iron (III) Oxalate?

$$\text{Fe}_2(\text{C}_2\text{O}_4)_3 = 375.6 \text{ g/mol}$$

4. How many moles are 500 grams Iron (III) Oxalate?

$$1.33 \text{ mol Fe}_2(\text{C}_2\text{O}_4)_3$$

5. How many grams is 1.78 moles Iron (II) Hydroxide?

$$\text{Fe}(\text{OH})_2 = 89.8 \text{ g/mol} \quad 159.8 \text{ g Fe}(\text{OH})_2$$

6. Calculate the volume of 3.44 moles methane, CH_4 , gas at STP:

$$77.06 \text{ liters CH}_4$$

7. What is the density of Nitrogen Dioxide at STP?

$$\text{NO}_2 \quad 46 \text{ g/mol} \quad 2.05 \text{ g/l NO}_2$$

8. What is the volume of 3.24×10^{22} molecules of Chlorine gas at STP?

$$1.21 \text{ liters Cl}_2$$

9. How many H atoms and O atoms are in a 500 ml water bottle?

$$\text{H atoms} = 3.33 \times 10^{25}$$

$$\text{O atoms} = 1.67 \times 10^{25}$$

10. A dime is an alloy (a mixture of two solids) of 75% Copper (Cu) and 25% Nickel (Ni) and has a mass of 2.27 grams; how many Cu, Ni, and total atoms does a dime contain?

$$\text{Cu} = 1.61 \times 10^{22}$$

$$\text{Ni} = 5.80 \times 10^{21}$$

$$\text{total atoms} = 2.19 \times 10^{22}$$

Given a container with 50-ml of Ethanol ($\text{C}_2\text{H}_6\text{O}$); the density of Ethanol is 0.789g/ml

11. Ethanol's molar mass 46 g/mol

12. Moles of ethanol 1.166 moles Ethanol

13. Carbon atoms 1.4×10^{24} C atoms

Chapter 10 Chemical Quantities Practice Problems Answers

Howard Stephen Stoker



Chapter 10 Chemical Quantities Practice Problems Answers:

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