

Answer Key

Chapter 10

Practice Problems

1. 9.45×10^{24} molecules CO_2
2. 3.27×10^{22} molecules H_2O
3. 1.53 mol Fe
4. 0.0341 mol sucrose
5. 1.08×10^{21} formula units CuSO_4
6. 839 g Sb
7. 5.53 g Se
8. 6.95 mol S
9. 12.5 mol He
10. 0.0213 mol Ni
11. 6.17×10^{21} atoms Pt
12. 1.88×10^{28} atoms S
13. 39.7 g Hg
14. 6.35×10^{-3} g
15. 4.36×10^{16} Pb ions
16. 0.220 mol C_6H_6
17. 2.856 mol KClO_3
18. 40.2 g SnSO_4
19. 3.40×10^2 g (340 g expressed to 3 significant figures)
20. 13.1 g Zn and 37.5 g $\text{Cu}(\text{NO}_3)_2$
21. 52.93% Al, 47.07% O
22. 16.39% Mg, 18.89% N, 64.72% O
23. 39.17% O
24. 49.30% N
25. The percent compositions will be the same because the ratio of carbon atoms to hydrogen atoms is the same, 1:1.
26. CH_2O
27. ZnN_2O_6
28. Al_4C_3
29. $\text{C}_3\text{H}_4\text{O}_3$
30. $\text{C}_8\text{H}_8\text{N}_2\text{O}_2$
31. $\text{B}_3\text{H}_6\text{N}_3$
32. $\text{Ag}_2\text{C}_2\text{O}_4$
33. P_2S_5 ; same as empirical formula
34. $\text{C}_9\text{H}_{12}\text{N}_6$
35. $\text{CeI}_3 \cdot 9\text{H}_2\text{O}$
36. $\text{Co}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$
37. $\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$
38. $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$

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John Muter



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