

1. **Strategize** between the basic units and volume. First, as a multiple of a solution and density, mass is required. This solution is given; the substance is a solution that is the general amount. It is also a substance that is dissolved in the solution more a solution. There is the solvent and solute in the solution is a solution solution.

2. What is the concentration of silver ion in each of the following aqueous solutions?

a. 0.20 M AgNO_3 $[\text{Ag}^+] = \frac{0.20 \text{ mol AgNO}_3}{1 \text{ L solution}} \times \frac{1 \text{ mol Ag}^+}{1 \text{ mol AgNO}_3} = 0.20 \text{ M}$

b. $0.20 \text{ M Ag}_2\text{SO}_4$ $[\text{Ag}^+] = \frac{0.20 \text{ mol Ag}_2\text{SO}_4}{1 \text{ L solution}} \times \frac{2 \text{ mol Ag}^+}{1 \text{ mol Ag}_2\text{SO}_4} = 0.40 \text{ M}$

c. $0.20 \text{ M Ag}_2\text{C}_2\text{O}_4$ $[\text{Ag}^+] = \frac{0.20 \text{ mol Ag}_2\text{C}_2\text{O}_4}{1 \text{ L solution}} \times \frac{2 \text{ mol Ag}^+}{1 \text{ mol Ag}_2\text{C}_2\text{O}_4} = 0.40 \text{ M}$

3. A solution is prepared by dissolving 25.0 g of sodium sulfate in enough water to prepare 250 mL of solution.

- a. What is the molarity of sodium sulfate in the solution?

$$\text{M}_\text{Na}_2\text{SO}_4 \text{ of } \text{Na}_2\text{SO}_4 = \frac{2 \text{ mol Na}^+ + 25.0 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}^+ + 25.0 \text{ g Na}_2\text{SO}_4} = \frac{2 \text{ mol Na}^+ + 25.0 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}^+ + 25.0 \text{ g Na}_2\text{SO}_4} = 0.40 \text{ M}$$

$$\text{M}_\text{Na}_2\text{SO}_4 \text{ of } \text{Na}_2\text{SO}_4 = \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}^+} = 0.40 \text{ M}$$

$$\text{M}_\text{Na}_2\text{SO}_4 \text{ of } \text{Na}_2\text{SO}_4 = \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}^+} = 0.40 \text{ M}$$

- b. What are the molarities of the sodium and sulfate ions in the solution?

$$[\text{Na}^+] = \frac{0.40 \text{ mol Na}_2\text{SO}_4}{1 \text{ L solution}} \times \frac{2 \text{ mol Na}^+}{1 \text{ mol Na}_2\text{SO}_4} = 0.80 \text{ M}$$

$$[\text{SO}_4^{2-}] = \frac{0.40 \text{ mol Na}_2\text{SO}_4}{1 \text{ L solution}} \times \frac{1 \text{ mol SO}_4^{2-}}{1 \text{ mol Na}_2\text{SO}_4} = 0.40 \text{ M}$$

- c. How many moles of sodium ions are present in 17 mL of the solution?

$$17 \text{ mL} \times \frac{0.80 \text{ mol Na}^+}{1 \text{ L solution}} = 0.014 \text{ mol Na}^+$$

4. A solution of $\text{Ba}_3(\text{PO}_4)_2$ which has a volume of 75.0 mL, contains 0.0045 moles of precipitate ions. What is the molarity of the $\text{Ba}_3(\text{PO}_4)_2$ solution?

$$\frac{0.0045 \text{ mol Ba}^{2+}}{0.075 \text{ L}} = \frac{0.0045 \text{ mol Ba}^{2+}}{0.075 \text{ L}} = 0.060 \text{ M Ba}_3(\text{PO}_4)_2$$

5. What are the three steps for calculating molarities?

Subtract moles of reacting materials in the solution and placing the units provided in the solution.

6. Explain the hydrophobic effect.

The OH groups of negative compounds are removed and placed very poorly with water. Compounds with many OH groups interact in poorly with water (they are called hydrophobic). When a hydrophobic compound is placed in water, the water molecules interact in a network that connects with the compound. As a result, hydrophobic compounds are not soluble in water.

7. Explain whether each of the following substances might be more soluble in H_2O or C_6H_6 .

a. BaF_2 (aq) b. $\text{C}_6\text{H}_5\text{OH}$ (aq) c. NaCl (aq) d. $\text{C}_6\text{H}_5\text{OH}$ (aq)

Hydrogen bonding substances (H_2O and $\text{C}_6\text{H}_5\text{OH}$) and ionic substances (NaCl and $\text{C}_6\text{H}_5\text{OH}$) are more soluble in water than in a nonpolar solvent, while ionic substances ($\text{C}_6\text{H}_5\text{OH}$) are more soluble in nonpolar solvents.

Chapter 10 Solutions Baf3m Chatt

Christian G. Meyer



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