CHAPTER 10 The Mole

Section 1 Measuring Matter

page 320-324

Practice Problems

pages 323-324

 Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

-1.51×1026 atoms of Zn

 Calculate the number of molecules in 11.5 mol of water (H₂O).

= 6.92×10³⁴ molecules H₂O

 Silver nitrate (AgNO₁) is used to make several different silver halides used in photographic films.
 How many formula units of AgNO₃ are there in 3.25 mol AgNO₁?

3.25 mol AgNO,
$$\times \frac{6.02 \times 10^{10} \text{ formula units AgNO,}}{1 \text{ mol AgNO,}}$$

= 1.96 × 1014 formula units AgNO,

 Challenge Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule. Ox.

5.00 mol
$$O_2 \times \frac{6.02 \times 10^{21} \text{ molecules } O_2}{1 \text{ mol} O_2}$$

- 5. How many moles contain each of the following?
 - 5.75×10²⁴ atoms Al

$$5.75 \times 10^{14}$$
 atoms AI $\times \frac{\text{I mol AI}}{6.02 \times 10^{19}}$ atoms AI = 9.55 mol AI

b. 2.50×10²⁰ atoms Fe

Chemistry: Matter and Change

- Challenge Identify the representative particle for each formula and convert the given number of representative particles to moles.
 - a. 3.75×10²⁴ CO₂

The representative particle is a molecule.

3.75×10²⁴ molecules CO₂

 $\times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$

b. 3.58×10²³ ZnCh

The representative particle is a formula unit.

3.58×1021 formula units ZnCl,

× 1 mol ZnCl, 6.02×10¹³ formula units ZnCl,

0.595 mol ZnCl₂

Section 1 Review

page 324

Explain why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

 State the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number (6.02×10²³) of representative particles.

Chapter 10 Study Guide The Mole Answer Key

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Table of Contents Chapter 10 Study Guide The Mole Answer Key

- 1. Understanding the eBook Chapter 10 Study Guide The Mole Answer Key
 - The Rise of Digital Reading Chapter 10 Study Guide The Mole Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 10 Study Guide The Mole Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Study Guide The Mole Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 10 Study Guide The Mole Answer Key
 - Personalized Recommendations
 - $\circ\,$ Chapter 10 Study Guide The Mole Answer Key User Reviews and Ratings
 - Chapter 10 Study Guide The Mole Answer Key and Bestseller Lists
- 5. Accessing Chapter 10 Study Guide The Mole Answer Key Free and Paid eBooks
 - Chapter 10 Study Guide The Mole Answer Key Public Domain eBooks
 - Chapter 10 Study Guide The Mole Answer Key eBook Subscription Services
 - Chapter 10 Study Guide The Mole Answer Key Budget-Friendly Options

- 6. Navigating Chapter 10 Study Guide The Mole Answer Key eBook Formats
 - o ePub, PDF, MOBI, and More
 - Chapter 10 Study Guide The Mole Answer Key Compatibility with Devices
 - Chapter 10 Study Guide The Mole Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 10 Study Guide The Mole Answer Key
 - Highlighting and Note-Taking Chapter 10 Study Guide The Mole Answer Key
 - Interactive Elements Chapter 10 Study Guide The Mole Answer Key
- 8. Staying Engaged with Chapter 10 Study Guide The Mole Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Study Guide The Mole Answer Key
- 9. Balancing eBooks and Physical Books Chapter 10 Study Guide The Mole Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Study Guide The Mole Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 10 Study Guide The Mole Answer Key
 - Setting Reading Goals Chapter 10 Study Guide The Mole Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 10 Study Guide The Mole Answer Key
 - Fact-Checking eBook Content of Chapter 10 Study Guide The Mole Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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