

Economics

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(Introductory Macroeconomics) (Chapter – 6) (Open Economy Macroeconomics)

(Class – XII)

Answer 15:

The rupee-dollar exchange rate is given as ₹30 in the year 2010. This implies that the price of a good (say, book) is \$1 in USA and ₹30 in India. It is given that prices have doubled in India over 20 years while they have remained fixed in the USA, so the cost of the book will be ₹60 in 2030 and the cost of the book remains the same i.e. \$1 in the USA since the prices are not changing. For these two prices to be equivalent, ₹60 must be worth of \$1. The rupee, therefore, will depreciate.

Question 16:

If inflation is higher in country A than in Country B, and the exchange rate between the two countries is fixed, what is likely to happen to the trade balance between the two countries?

Answer 16:

Country A has a higher inflation than country B. Since, the exchange rate is fixed, it is advantageous for country B to export goods to country A. Similarly, it is advantageous for country A to import goods from country B. On the other hand, it would be expensive for country A to export goods to country B. Thus, country A will have trade deficit as it will import more goods as compared to exports, from country B. Country B will import less goods as compared to exports, from country A. Hence, there is a trade surplus in country B.

Question 17:

Should a current account deficit be a cause for alarm? Explain.

Answer 17:

Current account deficit is the excess of total imports of goods, services and transfers over total exports of goods, services and transfers. This situation makes a country debtor to the rest of the world. But, this cannot be always treated as a cause for alarm because countries might be running in deficits (current account) to increase productivity and exports in future. Also, more investment will help in building capital stock, which in future will lead to rise in output.

Question 18:

Suppose $C = 100 + 0.75Y$, $D, I = -500$, $G = 750$, taxes are 20 per cent of income, $X = 150$, $M = 100 + 0.2Y$. Calculate equilibrium income, the budget deficit or surplus and the trade deficit or surplus.

Answer 18:

$$C = 100 + 0.75YD,$$

$$I = 500,$$

$$G = 750,$$

$$X = 150,$$

$$M = 100 + 0.2Y$$

$$\text{Equilibrium income (Y)} = C + c(Y - T) + I + G + X - M - mY$$

$$\Rightarrow Y = 100 + .75\left(Y - \frac{20}{100}Y\right) + 500 + 750 + 150 - 100 - 0.2Y$$

$$\Rightarrow Y = 1400 + \frac{75}{100} \times \frac{4Y}{5} - 0.2Y \Rightarrow Y = 1400 + \frac{3}{5}Y - 0.2Y \Rightarrow \frac{6Y}{10} = 1400$$

$$\Rightarrow Y = \frac{1400 \times 10}{6} = \frac{7000}{3}$$

$$\text{Government expenditure} = 750$$

$$\text{Government receipts (taxes)} = \frac{20}{100} \times \frac{7000}{3} = \frac{1400}{3} = 466.6$$

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**Joyce Oldham Appleby, Alan
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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- + 2\text{Ag}^+$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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may 11 2017 in the net ionic equation any ions that do not participate in the reaction called spectator ions are excluded as a result the net ionic equation shows only the species that

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1 $\text{K}_3\text{PO}_4 + \text{Al}(\text{OH})_3 \rightarrow \text{AlPO}_4 + 3\text{H}_2\text{O}$ ionic $3\text{PO}_4^{3-} + \text{Al}^{3+} \rightarrow \text{AlPO}_4 + 3\text{H}_2\text{O}$ net ionic $\text{PO}_4^{3-} + \text{Al}^{3+} \rightarrow \text{AlPO}_4$

aq 3 mg s

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oct 2 2019 the net ionic equation is a chemical equation for a reaction that lists only those species participating in the reaction the net ionic equation is commonly used in acid base

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jun 23 2022 what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic

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aug 10 2023 write the full ionic and net ionic equations for this reaction solution the elemental metals magnesium on the reactant side copper on the product side are neutral

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the first step to writing a net ionic equation is balancing the chemical equation present let s use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction notice that in writing the net ionic

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what remains when the spectator ions are removed is called the net ionic equation which represents the actual chemical change occurring between the ionic compounds cl aq

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh₄ aq cr₂o₇²⁻ aq ag₂cr₂o₇ s 2nh₄ aq

2f aq because two nh 4 aq

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