Net Ionic Equations

Balance the following equations and show their ionic & net ionic forms.

1) $AgNO_3(aq.)$ + KCI(aq.) \longrightarrow AgCI(s) + $KNO_3(aq.)$

lonic:

Net ionic:

2) $Ca(OH)_2(aq.) + Al_2(SO_4)_3(aq.) \longrightarrow Al(OH)_3(s) + CaSO_4(aq.)$

lonic:

Net ionic:

3) $NaCl(aq.) + Pb(NO_3)_2(aq.) \rightarrow PbCl_2(s) + NaNO_3(aq.)$

lonic:

Net ionic:

4) $Na_2CO_3(aq.)$ + $FeCl_2(aq.)$ \longrightarrow $FeCO_3(s)$ + NaCl(aq.)

lonic:

Net ionic:

5) $Mg(NO_3)_2(aq.) + Na_2CrO_4(aq.) \rightarrow NaNO_3(aq.) + MgCrO_4(s)$

Ionic:

Net ionic:

6) LiOH(aq.) + $VCI_3(aq.)$ \longrightarrow $V(OH)_3(s)$ + LiCI(aq.)

lonic:

Net ionic:

7) $BaBr_2(aq.) + Na_2SO_4(aq.) \longrightarrow BaSO_4(s) + NaBr(aq.)$

lonic:

Net ionic:

8) $Ni(NO_3)_2(aq.) + NaOH(aq.) \rightarrow Ni(OH)_2(s) + NaNO_3(aq.)$

lonic:

Net ionic:

Balancing Ionic Equation Practice

Lyman Churchill Newell

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