Balancing and Classifying Chemical Equations

<u>Directions</u>: First, balance each of the chemical equations below. Then, classify each reaction as synthesis, decomposition, single-replacement, double-replacement or combustion.

Balance the equation...

- 2 Sb + 3 Cl₂ → 2 SbCl₂
- 2. Mg + O₂ → 2 MgO
- CaCl₂ → ___ Ca + ___ Cl₂
- 2 NaClO₃ → 2 NaCl + 3 O₂
- Fe + 2 HCl → FeCl₂ + H₂
- CuO + ___ H₂ → ___ Cu + ___ H₂O
- 2 Al + 3 H₂SO₁ → Al₂(SO₁)₃ + 3 H₂
- MgBr₂ + Cl₂ → MgCl₂ + Br₂
- 9. ____SnO₂ + _2 C → ____Sn + _2 CO
- Pb(NO₃)₂ + ____ PbS + ___ PbS + ___ HNO₃
- 11. __2_HgO → __2_Hg + ___O₂
- 12. __2 KClO₃ → __2 KCl + __3 O₂
- 13. N₂ + 3 H₂ → 2 NH₂
- 14. 2 NaBr + Cl₂ → 2 NaCl + Br₂
- 15. Zn + 2 AgNO₃ → Zn(NO₃)₂ + 2 Ag
- 16. Sn + 2 Cl₂ → SnCl₄
- 17. ____Ba(OH)₂ → ____BaO + ____H₂O
- Mg(OH)₂ + __2 HCl → ___ MgCl₂ + __2 H₂O
- 19. ____Na₀CO₃ + __2_HCl → __2_NaCl + ____H₂CO₃
- 20. $NH_1NO_2 \rightarrow N_2 + 2H_2O$
- 21. 2 N₂ + 5 O₂ → 2 N₂O₃
- 22. MgCO₂ → MgO + CO₂
- 23. 2 KBr + Cl₂ → 2 KCl + Br₂
- 24. ____Zn + ____CuSO₄ → ____Cu + ____ZnSO₄
- 25. 4 P + 3 O₂ → P₆O₆

...and classify it.

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Balancing And Classifying Chemical Equations

Charles H. Corwin

Balancing And Classifying Chemical Equations:

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