

# Balancing Redox Equations

Equations for simple redox reactions can be balanced by inspection. Most redox equations, however, require more systematic methods. The equation-balancing process requires the use of oxidation numbers. In a balanced equation, both charge and mass are conserved. Although oxidation and reduction half-reactions occur together, their reaction equations are balanced separately and then combined to give the balanced redox-reaction equation.

## MAIN IDEA

**Oxidation and reduction reactions are balanced separately and then added together.**

The half-reaction method, or ion-electron method, for balancing redox equations consists of seven steps. Oxidation numbers are assigned to all atoms and polyatomic ions to determine which species are part of the redox process. The oxidation and reduction equations are balanced separately for mass and charge. They are then added together to produce a complete balanced equation. These seven steps are applied to balance the reaction of hydrogen sulfide and nitric acid. Sulfuric acid, nitrogen dioxide, and water are the products of the reaction.

1. Write the formula equation if it is not given in the problem. Then write the ionic equation.



2. Assign oxidation numbers. Delete substances containing only elements that do not change oxidation state.



The sulfur changes oxidation state from  $-2$  to  $+6$ . The nitrogen changes oxidation state from  $+5$  to  $+4$ . The other substances are deleted.



The remaining species are used in step 3.

## SECTION 2

### Main Idea

- ▶ Oxidation and reduction reactions are balanced separately and then added together.

## SECTION 2

### Plan and Prepare ▼

#### Classroom Catalyst

Students are already used to balancing equations by inspection and may resist the idea of following a procedure involving half-reactions. To emphasize the need for a systematic approach, write the equation for the reaction in the Spontaneous Redox Reaction demonstration (at the beginning of this chapter) on the board, and ask students to balance it.



The likely result will be



The equation is balanced as to mass but not balanced as to charge. Having made the point, use the half-reaction method to balance the equation.

The two half-reactions are



and



### Teach ▼

#### Teaching Tip

Remind students that in redox reactions, the number of electrons lost in the oxidation half of the reaction should equal the number of electrons gained in the reduction half of the reaction. The net charge and number of ions must be the same on both sides of the balanced equation.

## Differentiated Instruction

### INCLUSION

Learning-disabled and attention-deficit-disorder students can develop a study aid for understanding oxidation and reduction. Have them fold a piece of paper in half vertically. Label one side Oxidation and the other side Reduction. Students can fill each column with examples and descriptions of the appropriate change. For example, under Oxidation, they can write "loss of electrons," "increase in



# Balancing Equations Section 2

**Ernst Brauer, Franz Lawaczeck**



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Webster, Hermon Waldo Farwell, Elmer Reginald Drew, 1923      **The School World**, 1907      **Chemical Engineering**  
**Computing** American Institute of Chemical Engineers, 1972      A Mechanical Text-book; Or, Introduction to the Study of  
Mechanics and Engineering, Etc William John Macquorn Rankine, 1873      **Northwest Science**, 1927      *Supplement to*  
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### **Table of Contents Balancing Equations Section 2**

1. Understanding the eBook Balancing Equations Section 2
  - The Rise of Digital Reading Balancing Equations Section 2
  - Advantages of eBooks Over Traditional Books
2. Identifying Balancing Equations Section 2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Balancing Equations Section 2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Balancing Equations Section 2
  - Personalized Recommendations

- Balancing Equations Section 2 User Reviews and Ratings
- Balancing Equations Section 2 and Bestseller Lists
- 5. Accessing Balancing Equations Section 2 Free and Paid eBooks
  - Balancing Equations Section 2 Public Domain eBooks
  - Balancing Equations Section 2 eBook Subscription Services
  - Balancing Equations Section 2 Budget-Friendly Options
- 6. Navigating Balancing Equations Section 2 eBook Formats
  - ePub, PDF, MOBI, and More
  - Balancing Equations Section 2 Compatibility with Devices
  - Balancing Equations Section 2 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Balancing Equations Section 2
  - Highlighting and Note-Taking Balancing Equations Section 2
  - Interactive Elements Balancing Equations Section 2
- 8. Staying Engaged with Balancing Equations Section 2
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Balancing Equations Section 2
- 9. Balancing eBooks and Physical Books Balancing Equations Section 2
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Balancing Equations Section 2
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Balancing Equations Section 2
  - Setting Reading Goals Balancing Equations Section 2
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Balancing Equations Section 2
  - Fact-Checking eBook Content of Balancing Equations Section 2

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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