

NUCLEAR CHEMISTRY - WORKSHEET

1. When stable, nonradioactive nitrogen in the air is subject to neutron bombardment during a hydrogen bomb explosion, a single neutron is absorbed by the nitrogen nucleus and a proton is emitted (given off). Write a balanced nuclear equation for this reaction and name the other product of the reaction.



2. A radioactive element, A, with a mass of 212 and atomic number 84, emits an alpha particle and changes into element B. Element B emits a beta particle and is converted to element C.

a) Write nuclear equations for the formation of elements B and C.



b) What are the atomic masses and atomic number of elements A, B and C?

Masses are on top; atomic numbers at bottom.

c) Name elements A, B and C.

Polonium, Lead, and Bismuth.

3. Complete the following nuclear equations and state the general type:



a) ${}^0_{-1}\text{e}$ or ${}^0_0\gamma$	b) ${}^{210}_{81}\text{Tl}$
c) ${}^{129}_{50}\text{Sn}$	d) ${}^{182}_{70}\text{Yb}$
e) ${}^1_1\text{H}$	f) ${}^1_0\text{n}$

4. Write a balanced nuclear equation for each of the following:



Concept Review Nuclear Chemistry Answer Key

Gary S. Thorpe



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