

# Chemistry Study Guide

### Periodic Table of the Elements

<div><div>H</div><div>1</div><div>1.008</div></div>	<div><div>He</div><div>2</div><div>4.003</div></div>																	<div><div>He</div><div>2</div><div>4.003</div></div>															
<div><div>Li</div><div>3</div><div>6.941</div></div>	<div><div>Be</div><div>4</div><div>9.012</div></div>																	<div><div>B</div><div>5</div><div>10.811</div></div>	<div><div>C</div><div>6</div><div>12.011</div></div>	<div><div>N</div><div>7</div><div>14.007</div></div>	<div><div>O</div><div>8</div><div>15.999</div></div>	<div><div>F</div><div>9</div><div>18.998</div></div>	<div><div>Ne</div><div>10</div><div>20.180</div></div>										
<div><div>Na</div><div>11</div><div>22.990</div></div>	<div><div>Mg</div><div>12</div><div>24.305</div></div>																	<div><div>Al</div><div>13</div><div>26.982</div></div>	<div><div>Si</div><div>14</div><div>28.086</div></div>	<div><div>P</div><div>15</div><div>30.974</div></div>	<div><div>S</div><div>16</div><div>32.065</div></div>	<div><div>Cl</div><div>17</div><div>35.453</div></div>	<div><div>Ar</div><div>18</div><div>39.948</div></div>										
<div><div>K</div><div>19</div><div>39.098</div></div>	<div><div>Ca</div><div>20</div><div>40.078</div></div>	<div><div>Sc</div><div>21</div><div>44.956</div></div>	<div><div>Ti</div><div>22</div><div>47.883</div></div>	<div><div>V</div><div>23</div><div>50.942</div></div>	<div><div>Cr</div><div>24</div><div>51.996</div></div>	<div><div>Mn</div><div>25</div><div>54.938</div></div>	<div><div>Fe</div><div>26</div><div>55.845</div></div>	<div><div>Co</div><div>27</div><div>58.933</div></div>	<div><div>Ni</div><div>28</div><div>58.693</div></div>	<div><div>Cu</div><div>29</div><div>63.546</div></div>	<div><div>Zn</div><div>30</div><div>65.380</div></div>	<div><div>Ga</div><div>31</div><div>69.723</div></div>	<div><div>Ge</div><div>32</div><div>72.630</div></div>	<div><div>As</div><div>33</div><div>74.922</div></div>	<div><div>Se</div><div>34</div><div>78.960</div></div>	<div><div>Br</div><div>35</div><div>79.904</div></div>	<div><div>Kr</div><div>36</div><div>83.801</div></div>																
<div><div>Rb</div><div>37</div><div>85.468</div></div>	<div><div>Sr</div><div>38</div><div>87.62</div></div>	<div><div>Y</div><div>39</div><div>88.906</div></div>	<div><div>Zr</div><div>40</div><div>91.224</div></div>	<div><div>Nb</div><div>41</div><div>92.906</div></div>	<div><div>Mo</div><div>42</div><div>95.94</div></div>	<div><div>Tc</div><div>43</div><div>98.906</div></div>	<div><div>Ru</div><div>44</div><div>101.07</div></div>	<div><div>Rh</div><div>45</div><div>102.91</div></div>	<div><div>Pd</div><div>46</div><div>106.42</div></div>	<div><div>Ag</div><div>47</div><div>107.87</div></div>	<div><div>Cd</div><div>48</div><div>112.41</div></div>	<div><div>In</div><div>49</div><div>114.82</div></div>	<div><div>Sn</div><div>50</div><div>118.71</div></div>	<div><div>Sb</div><div>51</div><div>121.76</div></div>	<div><div>Te</div><div>52</div><div>127.6</div></div>	<div><div>I</div><div>53</div><div>126.91</div></div>	<div><div>Xe</div><div>54</div><div>131.29</div></div>																
<div><div>Cs</div><div>55</div><div>132.91</div></div>	<div><div>Ba</div><div>56</div><div>137.33</div></div>																	<div><div>Tl</div><div>81</div><div>204.38</div></div>	<div><div>Pb</div><div>82</div><div>207.2</div></div>	<div><div>Bi</div><div>83</div><div>208.98</div></div>	<div><div>Po</div><div>84</div><div>209</div></div>	<div><div>At</div><div>85</div><div>210</div></div>	<div><div>Rn</div><div>86</div><div>222</div></div>										
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		<div><div>La</div><div>57</div><div>138.91</div></div>																<div><div>Ce</div><div>58</div><div>140.12</div></div>	<div><div>Pr</div><div>59</div><div>140.91</div></div>	<div><div>Nd</div><div>60</div><div>144.24</div></div>	<div><div>Pm</div><div>61</div><div>144.91</div></div>	<div><div>Sm</div><div>62</div><div>150.36</div></div>	<div><div>Eu</div><div>63</div><div>151.96</div></div>	<div><div>Gd</div><div>64</div><div>157.25</div></div>	<div><div>Tb</div><div>65</div><div>158.93</div></div>	<div><div>Dy</div><div>66</div><div>162.50</div></div>	<div><div>Ho</div><div>67</div><div>164.93</div></div>	<div><div>Er</div><div>68</div><div>167.26</div></div>	<div><div>Tm</div><div>69</div><div>168.93</div></div>	<div><div>Yb</div><div>70</div><div>173.05</div></div>	<div><div>Lu</div><div>71</div><div>174.97</div></div>		
		<div><div>Ac</div><div>89</div><div>227</div></div>																<div><div>Th</div><div>90</div><div>232.04</div></div>	<div><div>Pa</div><div>91</div><div>231.04</div></div>	<div><div>U</div><div>92</div><div>238.03</div></div>	<div><div>Np</div><div>93</div><div>237</div></div>	<div><div>Pu</div><div>94</div><div>244</div></div>	<div><div>Am</div><div>95</div><div>243</div></div>	<div><div>Cm</div><div>96</div><div>247</div></div>	<div><div>Bk</div><div>97</div><div>247</div></div>	<div><div>Cf</div><div>98</div><div>251</div></div>	<div><div>Es</div><div>99</div><div>252</div></div>	<div><div>Fm</div><div>100</div><div>257</div></div>	<div><div>Md</div><div>101</div><div>258</div></div>	<div><div>No</div><div>102</div><div>259</div></div>	<div><div>Lr</div><div>103</div><div>262</div></div>		

## Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

## Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

## Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons ( $e^-$ )
- **Inner part** → *nucleus* → houses the protons ( $p^+$ ) and neutrons ( $n$ )
- Outer & Inner part → together these are the **subatomic particles**

## Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	$p^+$	1	+1	Nucleus
Neutron	$n$	1	0	Nucleus
Electron	$e^-$	$\sim 0$	-1	Orbit/Shell

# Conceptual Chemistry Study Guide

**Leo J. Malone**



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