

Flame Test Lab Report

Introduction:

In this laboratory experiment, we became Robert Bunsen and utilized qualitative techniques to identify different elements present in salt solutions by observing flame colors. The procedure involved adding drops of the solution to an iron wire and heating it over a Bunsen burner flame. The heat generated by the flame is crucial for inducing light radiation, as it imparts energy to the electrons in the metal, causing them to move to higher energy levels farther from the nucleus. However, these excited electrons are strongly attracted to exist closer to the nucleus in a more stable state. Consequently, the excited electrons eventually return to a lower energy level, releasing energy in the form of a photon. This emitted photon possesses a specific wavelength, and if it falls within the visible light spectrum, it can be observed as a flame. Given that each element has a unique electron configuration, the color of the flames can be used to identify elements, although some elements may emit very similar flames. At times, solutions contaminated with sodium may be dominated by the element's bright yellow flame. In such cases, the flame is observed through cobalt blue glass, which absorbs the yellow flame, allowing the color of the desired elements to be seen. Similar interference may occur in certain mixtures, where colors emitted by different elements in the mixture may blend. Consequently, the flame test is not entirely effective for identifying mixtures, but it proves useful for determining specific elements.

Procedure:

The procedure was the same as in the lab manual.

Data Analysis:

| Solution | Flame Color | Identification of Element |
|----------|-----------------------|---------------------------|
| 1 | Bright Orange | Li |
| 2 | Light Blue | Pb |
| 3 | Bluish White | As |
| 4 | Bright White | Mg |
| 5 | Orange Red | Ca |
| 6 | Green | Cu |
| 7 | Yellow-Green | Ba |
| 8 | Yellow-Orange | Na |
| 9 | Red with cobalt glass | K |
| 10 | Vivid Red | Sr |

Conclusion Of Flame Test Lab

Steven Murov, Brian Stedjee



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