

Now _____ Date _____ Class _____

CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Covalent Bonding

Section 9.1 The Covalent Bond

In your textbook, read about the nature of covalent bonds.

Use each of the terms below just once to complete the passage.

covalent bond	molecule	sigma bond	exothermic	p bond
---------------	----------	------------	------------	--------

When sharing of electron occurs, the attachment between atoms that results is called a(n) **(1) covalent bond**. When such an attachment is formed, bond dissociation energy is released, and the process is **(2) exothermic**. When two or more atoms bond by means of electron sharing, the resulting particle is called a(n) **(3) molecule**. If the electrons shared are contained between the two atoms, the attachment is called a(n) **(4) sigma bond**. If the sharing involves the overlap of parallel orbitals, the attachment is called a(n) **(5) p bond**.

In your textbook, read about single and multiple bonds and bond strengths.

Circle the letter of the choice that best completes the statement or answers the question.

- Is what form the elements such as hydrogen, nitrogen, and oxygen normally exist?
 - as single atoms
 - (b) as molecules containing two atoms**
 - as molecules containing three atoms
 - as molecules containing four atoms
- How many electrons are shared in a double covalent bond?
 - none
 - one
 - two
 - (d) four**
- Bond length is the distance between
 - two molecules of the same substance
 - (b) the nucleus of two attached atoms**
 - the electron in two attached atoms
 - the orbitals of two attached atoms
- Which of the following relationships relating to bond length is generally correct?
 - (a) the shorter the bond, the stronger the bond**
 - the shorter the bond, the weaker the bond
 - the shorter the bond, the lower the electron in it
 - the shorter the bond, the lower the bond dissociation energy

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49

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CHAPTER 9 STUDY GUIDE FOR CONTENT MASTERY

Section 9.2 Naming Molecules

In your textbook, read about how binary compounds and acids are named from their formulas.

For each statement below, write true or false.

- Binary molecular compounds are generally composed of a metal and a nonmetal. **false**
- The second element in the formula of a binary compound is named using the suffix -ite. **false**
- The prefix zero indicates three atoms. **false**
- The prefix four indicates six atoms. **true**
- In naming the first element in a formula, the prefix zero is not used. **true**
- For binary acids, the hydrogen part of the compound is named using the prefix hydro. **true**
- An oxyacid contains only two elements. **false**
- If the name of the anion of an oxyacid ends in -ite, the acid name contains the suffix -ous. **false**

In your textbook, read about naming molecular compounds and oxoacids.

For each item in Column A, write the letter of the matching item in Column B.

Column A

- (c)** N_2O
- (d)** H_2CO_3
- (g)** H_2SO_4
- (h)** N_2O_5
- (f)** HNO_3
- (i)** HNO_2
- (b)** HBr
- (e)** HBrO_3

Column B

- hydrobromic acid
- nitrogen trioxide
- carbon monoxide
- nitrous acid
- nitric acid
- nitrous acid
- nitric acid
- carbon dioxide

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1. Understanding the eBook Content Mastery Chapter 1 answers
 - The Rise of Digital Reading Content Mastery Chapter 1 answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Content Mastery Chapter 1 answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Content Mastery Chapter 1 answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Content Mastery Chapter 1 answers
 - Personalized Recommendations
 - Content Mastery Chapter 1 answers User Reviews and Ratings
 - Content Mastery Chapter 1 answers and Bestseller Lists
5. Accessing Content Mastery Chapter 1 answers Free and Paid eBooks
 - Content Mastery Chapter 1 answers Public Domain eBooks
 - Content Mastery Chapter 1 answers eBook Subscription Services
 - Content Mastery Chapter 1 answers Budget-Friendly Options
6. Navigating Content Mastery Chapter 1 answers eBook Formats

- ePub, PDF, MOBI, and More
 - Content Mastery Chapter 1 answers Compatibility with Devices
 - Content Mastery Chapter 1 answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Content Mastery Chapter 1 answers
 - Highlighting and Note-Taking Content Mastery Chapter 1 answers
 - Interactive Elements Content Mastery Chapter 1 answers
 8. Staying Engaged with Content Mastery Chapter 1 answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Content Mastery Chapter 1 answers
 9. Balancing eBooks and Physical Books Content Mastery Chapter 1 answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Content Mastery Chapter 1 answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Content Mastery Chapter 1 answers
 - Setting Reading Goals Content Mastery Chapter 1 answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Content Mastery Chapter 1 answers
 - Fact-Checking eBook Content of Content Mastery Chapter 1 answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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