Objective 4: Students will know the different types of chemical reactions by looking at a chemical equation.

- 1. Explain what happens during each type of chemical reaction.
 - Synthesis- two or more compounds become one
 - Combustion- oxygen is a reactant, water and carbon dioxide are products
 - Decomposition- one compound becomes two or more
 - Single displacement- one switch occurs
 - Double displacement- two switches occur
- 2. State what type of chemical reaction each of the following equations show.
 - C₁₀H₈ + 12 O₂ ---> 10 CO₂ + 4 H₂O

combustion

8 Fe + S₈ ---> 8 FeS

synthesis

2 H,O ---> 2 H, + O,

....

decomposition

- Mg + 2 H₂O ---> Mg(OH)₂ + H₂
 single displacement
- Pb(NO₃)₂ + 2 KI ---> Pbl₂ + 2 KNO₃

double displacement

NaOH + KNO₃ --> NaNO₃ + KOH

double displacement

CH₄ + 2 O₂ --> CO₂ + 2 H₂O

combustion

2 Fe + 6 NaBr --> 2 FeBr₃ + 6 Na

single displacement

CaSO₄ + Mg(OH)₂ --> Ca(OH)₂ + MgSO₄

double displacement

- Pb + O₂ --> PbO₂
 synthesis
- Na₂CO₃ --> Na₂O + CO₂
 decomposition

Describing Chemical Reactions Answers

Antony C. Wilbraham, Dennis D. Staley, Michael S. Matta, Edward L. Waterman, Prentice-Hall Staff

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