

Chapter 6 Section 2: Describing Chemical Reactions

- I. What are chemical equations?
 - a. Is a short, _____ way to show a chemical _____, using _____ instead of words.
 - b. Uses chemical formulas with _____ and coefficients to _____ a reaction
 - i. A formula is a _____ of symbols that represents elements in the compound
 - ii. The substances at the _____ of the equation are called _____, when the reaction is complete, on the right you have _____.
 - iii. If you have _____ reactants or products they are separated by a + _____
 1. Reactant 1 + reactant 2 ----> Product 1 + product 2
 2. See figures 7 and 8, page 195
- II. Conversion of mass
 - a. During a chemical reaction, matter is _____ nor destroyed, it just changes _____.
 - b. All the atoms _____ at the _____ of the reaction are present at the _____
 - c. The principle of conservation of mass states that in a chemical reaction, the _____ of the reactants must _____ the total mass of the products
- III. Open and closed systems
 - a. Open: matter can _____ from or _____ to the surroundings.
 - i. Burning match
 - b. Closed: matter is _____ to enter or leave the system
 - i. Eg, reaction is sealed in a plastic bag
 - ii. See figure 10, page 197
- IV. Balancing Chemical Equations
 - a. To describe a reaction _____, a chemical equation must show the _____ of each type of _____ on both sides of the equation

Describing Chemical Reactions Answers Section 2

British Columbia. Dept. of Education



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