

Balancing Equations Worksheet – Answers

Note to students: It is acceptable to leave spaces blank when balancing equations – blank spaces are interpreted as containing the number "1".

- 1) $\text{Na}_3\text{PO}_4 + 3 \text{ KOH} \rightarrow 3 \text{ NaOH} + \text{K}_3\text{PO}_4$
- 2) $\text{MgF}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{ LiF}$
- 3) $\text{P}_4 + 3 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_3$
- 4) $2 \text{ RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + 2 \text{ RbF}$
- 5) $2 \text{ AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{ Ag}$
- 6) $\text{CF}_4 + 2 \text{ Br}_2 \rightarrow \text{CBr}_4 + 2 \text{ F}_2$
- 7) $2 \text{ HCN} + \text{CuSO}_4 \rightarrow \text{H}_2\text{SO}_4 + \text{Cu}(\text{CN})_2$
- 8) $\text{GaF}_3 + 3 \text{ Cs} \rightarrow 3 \text{ CsF} + \text{Ga}$
- 9) $\text{BaS} + \text{PtF}_2 \rightarrow \text{BaF}_2 + \text{PtS}$ (already balanced)
- 10) $\text{N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- 11) $2 \text{ NaF} + \text{Br}_2 \rightarrow 2 \text{ NaBr} + \text{F}_2$
- 12) $\text{Pb}(\text{OH})_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + \text{PbCl}_2$
- 13) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + \text{Al}_2(\text{SO}_4)_3$
- 14) $\text{CH}_4 + 2 \text{ O}_2 \rightarrow \text{CO}_2 + 2 \text{ H}_2\text{O}$
- 15) $2 \text{ Na}_3\text{PO}_4 + 3 \text{ CaCl}_2 \rightarrow 6 \text{ NaCl} + \text{Ca}_3(\text{PO}_4)_2$
- 16) $2 \text{ K} + \text{Cl}_2 \rightarrow 2 \text{ KCl}$
- 17) $2 \text{ Al} + 6 \text{ HCl} \rightarrow 3 \text{ H}_2 + 2 \text{ AlCl}_3$
- 18) $\text{N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
- 19) $\text{SO}_2 + 2 \text{ Li}_2\text{Se} \rightarrow \text{SSe}_2 + 2 \text{ Li}_2\text{O}$
- 20) $2 \text{ NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$

Balancing Equations 4 Answer Key

Ralph H. Petrucci



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Instructors Resource Manual John McMurry, Robert A. Pribush, Robert Fay, 2000-12-27

Student Solutions Manual for Whitten, Davis, Peck, and Stanley's General Chemistry, 7th Ed Wendy L. Keeney-Kennicutt, 2004 Contains answers and solutions to all even numbered end of chapter exercises Solutions are divided by section for easy reference by students

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General Chemistry Ralph H. Petrucci, 1989

Chemical Principles with Qualitative Analysis William L. Masterton, Emil J. Slowinski, Conrad L. Stanitski, 1986

Electrical Engineering, 1917

Chemical Reaction Engineering, 1972

Physical Science George A. Williams, Richard Barnes, Ray Doerhoff, Max Bolen, 1979 A physical science text stressing an awareness of the environment

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Proceedings ,1992 Transactions of the American Nuclear Society American Nuclear Society,1962 *Chemical*

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Physics, Enhanced Problems Version David Halliday,Robert Resnick,Jearl Walker,2003 A text for calculus based physics

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