

Balancing Equations Worksheet – Answers

Note to students: It is acceptable to leave spaces blank when balancing equations – blank spaces are interpreted as containing the number "1".

- 1) $\text{Na}_3\text{PO}_4 + 3 \text{ KOH} \rightarrow 3 \text{ NaOH} + \text{K}_3\text{PO}_4$
- 2) $\text{MgF}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2 \text{ LiF}$
- 3) $\text{P}_4 + 3 \text{ O}_2 \rightarrow 2 \text{ P}_2\text{O}_3$
- 4) $2 \text{ RbNO}_3 + \text{BeF}_2 \rightarrow \text{Be}(\text{NO}_3)_2 + 2 \text{ RbF}$
- 5) $2 \text{ AgNO}_3 + \text{Cu} \rightarrow \text{Cu}(\text{NO}_3)_2 + 2 \text{ Ag}$
- 6) $\text{CF}_4 + 2 \text{ Br}_2 \rightarrow \text{CBr}_4 + 2 \text{ F}_2$
- 7) $2 \text{ HCN} + \text{CuSO}_4 \rightarrow \text{H}_2\text{SO}_4 + \text{Cu}(\text{CN})_2$
- 8) $\text{GaF}_3 + 3 \text{ Cs} \rightarrow 3 \text{ CsF} + \text{Ga}$
- 9) $\text{BaS} + \text{PtF}_2 \rightarrow \text{BaF}_2 + \text{PtS}$ (already balanced)
- 10) $\text{N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$
- 11) $2 \text{ NaF} + \text{Br}_2 \rightarrow 2 \text{ NaBr} + \text{F}_2$
- 12) $\text{Pb}(\text{OH})_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + \text{PbCl}_2$
- 13) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + \text{Al}_2(\text{SO}_4)_3$
- 14) $\text{CH}_4 + 2 \text{ O}_2 \rightarrow \text{CO}_2 + 2 \text{ H}_2\text{O}$
- 15) $2 \text{ Na}_3\text{PO}_4 + 3 \text{ CaCl}_2 \rightarrow 6 \text{ NaCl} + \text{Ca}_3(\text{PO}_4)_2$
- 16) $2 \text{ K} + \text{Cl}_2 \rightarrow 2 \text{ KCl}$
- 17) $2 \text{ Al} + 6 \text{ HCl} \rightarrow 3 \text{ H}_2 + 2 \text{ AlCl}_3$
- 18) $\text{N}_2 + 3 \text{ F}_2 \rightarrow 2 \text{ NF}_3$
- 19) $\text{SO}_2 + 2 \text{ Li}_2\text{Se} \rightarrow \text{SSe}_2 + 2 \text{ Li}_2\text{O}$
- 20) $2 \text{ NH}_3 + \text{H}_2\text{SO}_4 \rightarrow (\text{NH}_4)_2\text{SO}_4$

Balancing Equations Activity Answer Key

James N. Jensen



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