

# Writing Chemical Formulas and Balancing Chemical Equations Webquest

*This activity goes with Chapter 9 in your textbook.*

## LET'S BEGIN.

Start by reading the following review of how to name compounds.

First let's review how to name compounds.

REMEMBER, if it is ionic (metal + nonmetal), you write the name of the metal first and end the nonmetal with -ide. If you are writing the formula then you have to make the cation and anion charges (also called oxidation numbers) equal zero. For example, potassium sulfide contains the metal cation K with a charge of +1 and the nonmetal anion S with a charge of -2. So we would need 2 +1 for every 1 -2 ( $+1+1-2=0$ ). The formula would be  $K_2S$ .

To figure out cation and anion charges look at the periodic table. Group 1 has a +1 charge, group 2 has a +2, group 13 (where Al is) has a +3, skip group 14 because they don't form ions, group 15 has a -3 charge, group 16 has a -2 charge, group 17 has a -1 charge, and group 18 doesn't react.

periodic table group	1 or alkali metals	2 or alkaline earth metals	13 or Aluminum	14 or Carbon Family	15 or Nitrogen Family	16 or Oxygen Family	17 or Halogens	18 or Noble Gases
charge	+1	+2	+3	Don't form ions	-3	-2	-1	Don't react

Transition metal (the short columns in the middle of the table) charges are given to you in the form of Roman numerals. (I) = +1, (II) = +2, (III) = +3, (IV) = +4.

Polyatomic ions are compounds that have a charge. You have your polyatomic cheat sheet to use whenever you need them.

polyatomic ion	formula	charge
carbonate	$CO_3$	-2
sulfate	$SO_4$	-2
ammonium	$NH_4$	+1
nitrate	$NO_3$	-3

For covalent compounds (nonmetal + nonmetal) you use prefixes to name the compound. The prefix stands for how many of that atom there are in the compound. You must know the prefixes for the test. For example, carbon dioxide

# Balancing Equations Webquest Answers

**Christoph Steiner**



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