

## Study Card: Solvation

### What Does Dissolving Involve?

When a solid dissolves in water, it's not *gone*. It has changed its state. As a solid, a solute particle (ion, molecule, etc.) is attracted to other solute particles. But when it dissolves in  $\text{H}_2\text{O}$ , it changes to the aqueous state and becomes attracted to the  $\text{H}_2\text{O}$  molecules.

### The Dissolving of Ionic Solids

Dissolving involves more than just *mixing*. When table salt ( $\text{NaCl}$ ) is added to water, the ions break off the crystal one-by-one and are surrounded by  $\text{H}_2\text{O}$  molecules. The O atom of  $\text{H}_2\text{O}$  surrounds the + ions. The H atoms of  $\text{H}_2\text{O}$  surround the - ions. This is known as **solvation**; it prevents the ions from returning to the crystal lattice. All ionic solids are dissolved in this manner.

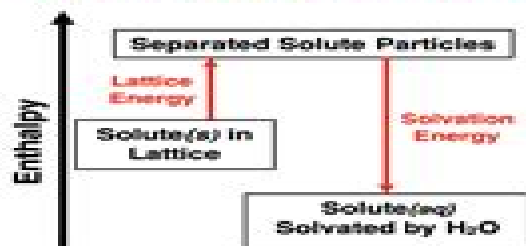
### The Dissolving of Molecular Solids

Molecular solids consist of molecules instead of ions. When table sugar (sucrose,  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ) dissolves, the molecules break off the crystal one-by-one and are surrounded by  $\text{H}_2\text{O}$  molecules. Hydrogen bonding forces hold the sucrose molecules together. So  $\text{H}_2\text{O}$  surrounds the -OH groups of a sucrose molecule, forming its own hydrogen bonds with sucrose and preventing the molecule from returning to the crystal. This is how **solvation** helps molecular solids dissolve.

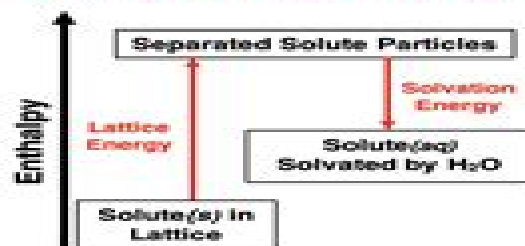
### An Energy Perspective on Dissolving

Dissolving can be thought of as involving two steps. First, solute particles are separated from the crystal. This requires energy; the amount is roughly equal to the **lattice energy**. Second, the  $\text{H}_2\text{O}$  molecules surround the solute particles. This releases energy; the amount is roughly equal to the **solvation energy**. Dissolution is more likely to occur if the solvation energy is greater than the lattice energy.

#### Exothermic Dissolution



#### Endothermic Dissolution



## The Dissolving Process



Solute particles attract each other.

Solute particles attract water.

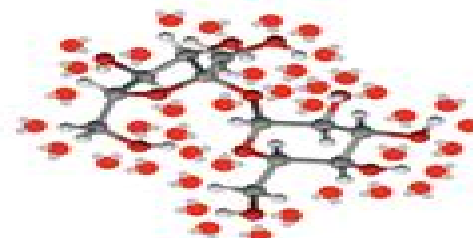
### Solvation

Polar  $\text{H}_2\text{O}$  molecules surrounding + and - ions.



### Solvation

$\text{H}_2\text{O}$  molecules attract and surround sucrose.



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