

The **concentration** of a **solution** tells you how much **solute** that is dissolved in a given amount of **solute** (**water**).

The **molarity** is the concentration of a solution.

$$\text{Molarity} = \frac{\text{Moles}}{\text{Litre}}$$

1. **2.255 moles** of NaCl is dissolved in **4.0 L** of water, calculate the molarity.

$$\text{Molarity} = \frac{2.255 \text{ moles}}{4.0 \text{ L}}$$

$$= 0.56 \text{ moles/litre}$$

$$= 0.56 \text{ M (Molar)}$$

where **M** means moles per 1 litre

# Concept Review Ii Concentration And Molarity

**Cambridge Philosophical Society**



## Concept Review Ii Concentration And Molarity:

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