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# Oxidation-Reduction Reactions

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  - (a) adjusting operficients to balance atoms.
  - (b) adjusting onefficients in electron equations to balance numbers of electrons lost and gained.
  - (c) adjusting subscripts to balance atoms.
  - (d) writing two separate electron equations.
- MinO<sub>1</sub> can be reclased to MinO<sub>2</sub>.
  - Assists the existation number to Ma in these two. minimum federal
  - b. How many electrons are named per Mn atom in their residue tions?
  - c. If 0.50 moll of MnO; is reduced, how many electrons are gained?
- 3. lodide ions can be oxidized to form iodine. Write the balanced oxidation halfreaction for the oxidation of indide to indise.
- Some bleaches contain accross chiorine as the active ingredient. Accross chloring is made by dissolving chloring gas in water. Aqueous chloring is: capable of exidising from(II) less to imm(III) lone. When from(II) less are oxidized, chloride loss are formed.
  - Write equations for the two half-reactions involved. Label them exidence or renductions.
  - Write the balanced ionic equation for the redox seaction between accesus. ethilosolime and incom/III).
  - Show that the equation in part b is balanced by charge.

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Fred M. Dewey

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