

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is air (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen in air is the (2) solvent and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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Chemistry: Matter and Change • Chapter 15

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**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

### Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

**The sodium and chloride ions are separated and surrounded by the water molecules.**

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

**Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

**The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**

13. List three ways that the rate of solvation may be increased.

**stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

**true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.

**false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

**false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

**true** 17. In a saturated solution, solvation and crystallization are in equilibrium.

**true** 18. Additional solute can be dissolved in an unsaturated solution.

**true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# **Content Mastery Study Guide Answers**

**Leslie S Treas, Judith M Wilkinson**



## **Content Mastery Study Guide Answers:**

**Study Guide to DSM-5-TR®** Laura Weiss Roberts, Alan K. Louie, Matthew L. Edwards, 2024-12-12 The ultimate companion volume to DSM 5 TR the Study Guide is designed to help clinical learners teachers and practitioners in psychiatry psychology and social work understand and apply diagnostic criteria Medical students residents and fellows and trainees throughout the clinical neurosciences will find this text to be invaluable in developing their diagnostic abilities Readers learn key clinical content through concise chapters and can assess their knowledge with more than 100 multiple choice questions The book is organized into three sections Part I delves into foundational concepts of psychiatric diagnoses and explores how diagnostic assessment fits into the process of working therapeutically with patients In this new edition Part I includes a new chapter examining structural and cultural considerations in arriving at a diagnosis Part II focuses on the diagnostic classes within DSM 5 TR and features clinical case vignettes with real life complexity relevant to psychiatric diagnoses diagnostic pearls that can be applied to daily clinical practice and a self assessment section with key concepts questions for further discussion patient cases and Q As Part III includes more than 100 multiple choice questions with an answer key These questions cover a range of diagnoses and patient presentations from across DSM 5 TR helping readers review their learning and assess their skills Taking a patient centered approach that complements the more disorder centered organization of DSM 5 TR this Study Guide provides context for DSM 5 TR s diagnoses while offering an array of interesting and memorable cases Rich in thought provoking detail and designed for repeated use this is an indispensable guide for learners educators and clinicians

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