

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Derive the molecular formula:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Given: Empirical formula =

Mass

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Given: Empirical formula = Mass of empirical formula =

Mass

3. A compound is found to contain 69.6% carbon, 5.8% hydrogen and 24.6% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Given: Carbon = 69.6%, Hydrogen = 5.8%, Oxygen = 24.6%

Given: Molecular mass = 174 g/mol

Empirical formula =

Mass

4. A compound is found to contain 69.6% carbon, 5.8% hydrogen and 24.6% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Given: Carbon = 69.6%, Hydrogen = 5.8%, Oxygen = 24.6%

Given: Molecular mass = 174 g/mol

Empirical formula =

Mass

5. A compound is found to contain 69.6% carbon, 5.8% hydrogen and 24.6% oxygen. Its molecular mass is 174 g/mol. What is its molecular formula?

Given: Carbon = 69.6%, Hydrogen = 5.8%, Oxygen = 24.6%

Given: Molecular mass = 174 g/mol

Empirical formula =

Mass

Determining Molecular Formula True Formulas Answer Key

**Commission on Undergraduate
Education in the Biological Sciences**



Determining Molecular Formula True Formulas Answer Key:

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Preparatory Chemistry Howard Stephen Stoker, 1985

The Development of Chemical Principles Cooper Harold Langford, Ralph Alonzo

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