

B = Bonding & Molecular Structure

Table B.1 • Some Average Single and Multiple Bond Energies (kJ/mol)

	H	C	N	O	F	Cl	S	Br	I	At
H-H	436									
C-H		413								
C-C		347								
C-N		305	391							
C-O		359	352	544						
C-F		485	483	517	453					
C-Cl		339	339	327	327	242				
C-Br		276	276	248	248	208				
C-I		238	238	213	213	172				
C-At										
N-H			391							
N-C			305							
N-N			160							
N-O			201							
O-H				467						
O-C				359						
O-N				201						
O-O				146						
S-H							347			
S-C							259			
S-S							226			
S-Br							192			
S-I							170			
S-At										
F-H	436				453					
F-C		485			453					
F-N			483		453					
F-O				517	453					
F-Cl					327					
F-Br					248					
F-I					213					
F-At										
Cl-H						242				
Cl-C						327				
Cl-N						327				
Cl-O						327				
Cl-S						208				
Cl-Br							192			
Cl-I							170			
Cl-At										
Br-H								238		
Br-C								276		
Br-N								276		
Br-O								248		
Br-S								208		
Br-I									172	
Br-At										
I-H									238	
I-C									238	
I-N									213	
I-O									213	
I-S									172	
I-Br										172
I-I										151
I-At										

Multiple Bonds			
C=C	614	C≡C	839
C=O	743	C≡O	1072
C=N	891	C≡N	960
C=O (gas, $D_{0,0}$)	1072	C≡O (gas, $D_{0,0}$)	1098

1. Write the balanced equation for the combustion of 1 mole of glucose, $C_6H_{12}O_6$, completely in O_2 .



2. Draw the structural formula showing Lewis structures for each of the molecules involved.

B = 1-17 **B = 18-19** **B = 20-21** **B = 22-23** **B = 24-25** **B = 26-27** **B = 28-29** **B = 30-31**



3. Calculate the bond energies of the reactants

$\Sigma C-H$ in $C_6H_{12}O_6$	1212	$\Sigma C-O$	3128
$\Sigma C=O$ in O_2	314	$\Sigma C=O$	374
$\Sigma O-H$ in H_2O	972	$\Sigma C=O$	1274

- Calculate the bond energies of the products

$\Sigma C=O$ in CO_2	2718	$\Sigma O-H$	374
$\Sigma C-O$ in H_2O	912	$\Sigma C=O$	2728

Total Bond Energy (Reactants) =	4667 kJ	Total Bond Energy (Products) =	5970 kJ
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4. ΔH_{comb} of Glucose Energy = _____, ΔH_{comb} of glucose = _____

What is the ΔH_{comb} based on combustion? = $-1507 kJ$

Bonding And Molecular Structure Worksheet Answers South Pasadena

Charles Alfred Coulson



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Chemical Structure and Bonding Roger L. DeKock, Harry B. Gray, 1989 Designed for use in inorganic physical and quantum chemistry courses this textbook includes numerous questions and problems at the end of each chapter and an Appendix with answers to most of the problems

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Structure and Bonding Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes

Chemical Bonding and the Geometry of Molecules George Eugene Ryschkewitsch, 1965

Bonding and Molecular Structure Russell S. Drago, 1971

The Shape and Structure of Molecules Charles Alfred Coulson, 1973

A Pictorial Approach to Molecular Bonding John G. Verkade, 2012-12-06

With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example

Molecules and the Chemical Bond Henry A. Bent, 2011

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Molecular Structure and Energetics; Volume 1: Chemical Bonding Models, 1986 **The Determination of Molecular Structure** Peter Jaffrey Wheatley, 1962

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