

## Chapter 1

### Some Basic Concepts of Chemistry

#### Question 1.1:

Calculate the molar mass of the following:

- (i)  $\text{H}_2\text{O}$                       (ii)  $\text{CO}_2$                       (iii)  $\text{CH}_4$

#### Answer 1.1:

- (i)  $\text{H}_2\text{O}$ :

The molecular mass of water,  $\text{H}_2\text{O}$   
 $= (2 \times \text{Atomic mass of hydrogen}) + (1 \times \text{Atomic mass of oxygen})$   
 $= [2(1.0084) + 1(16.00 \text{ u})]$   
 $= 2.016 \text{ u} + 16.00 \text{ u} = 18.016 = 18.02 \text{ u}$

- (ii)  $\text{CO}_2$ :

The molecular mass of carbon dioxide,  $\text{CO}_2$   
 $= (1 \times \text{Atomic mass of carbon}) + (2 \times \text{Atomic mass of oxygen})$   
 $= [1(12.011 \text{ u}) + 2(16.00 \text{ u})]$   
 $= 12.011 \text{ u} + 32.00 \text{ u} = 44.01 \text{ u}$

- (iii)  $\text{CH}_4$ :

The molecular mass of methane,  $\text{CH}_4$   
 $= (1 \times \text{Atomic mass of carbon}) + (4 \times \text{Atomic mass of hydrogen})$   
 $= [1(12.011 \text{ u}) + 4(1.008 \text{ u})]$   
 $= 12.011 \text{ u} + 4.032 \text{ u} = 16.043 \text{ u}$

#### Question 1.2:

Calculate the mass per cent of different elements present in sodium sulphate ( $\text{Na}_2\text{SO}_4$ ).

#### Answer 1.2:

The molecular formula of sodium sulphate is ( $\text{Na}_2\text{SO}_4$ ).

Molar mass of  $\text{Na}_2\text{SO}_4 = [(2 \times 23.0) + (32.066) + 4(16.00)] = 142.066 \text{ g}$

Mass percent of an element =  $\frac{\text{Mass of that element in the compound}}{\text{Molar Mass of the compound}} \times 100$

$\therefore$  Mass percent of sodium:  $\frac{46.0 \text{ g}}{142.066 \text{ g}} \times 100 = 32.379 = 32.4\%$

Mass percent of sulphur:  $\frac{32.066 \text{ g}}{142.066 \text{ g}} \times 100 = 22.57 = 22.6\%$

Mass percent of oxygen:  $\frac{64.0 \text{ g}}{142.066 \text{ g}} \times 100 = 45.049 = 45.05\%$

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**Victor M. Corman**



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