

COORDINATION COMPOUND

Introduction:-



Double salt

- Completely ionise in aq. medium.
- Lose their identity in aq. soln.
- Each ion gives positive test.
- Individual ions do not lose their own identity.

Examples

- ① Carnallite $[KCl \cdot MgCl_2 \cdot 6H_2O]$
- ② Mohar's salt $[FeSO_4(NH_4)_2SO_4 \cdot 6H_2O]$
- ③ Potash alum $[K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O]$

Complex salt

- Do not ionise completely.
- Retain their identity in aq. medium.
- Each ion do not give positive test.
- Ions lose their own identity.

Examples:- $K_4[Fe(CN)_6]$, $[Pt(NH_3)_2Cl_2]$ etc.

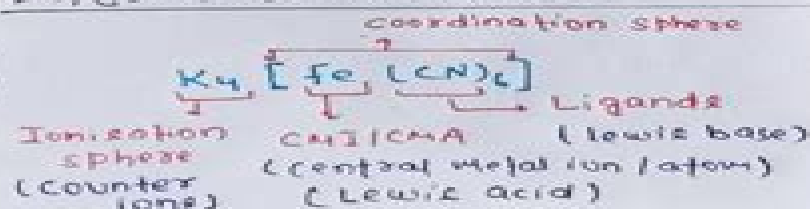
\hookrightarrow Complex salt \rightarrow σ coordinate bonds present b/w ligand (e^- pair donors) & Metal ion (e^- pair acceptor).

Q \rightarrow Why d-block element (Metal) form more no. of complexes?

Reason:- ① Availability of vacant orbital

② Small size & High ϕ nature

$\phi \uparrow$, Covalent nature of Bond \uparrow



Definitions

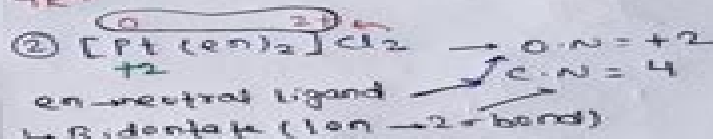
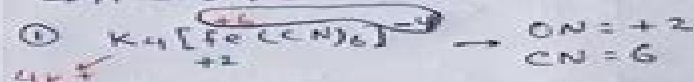
① Coordination number:- No. of ligand donor atoms to which metal is directly bonded.

or
 \hookrightarrow coordination no. is equal to σ -bonds present b/w ligand & CMI.

According to NCERT, π -bond b/w CMI & ligands will not be included in CN.

② oxidation no. of central atom:-

\hookrightarrow Andaz & Bahar ke charge ka hi saaf



③ Homoleptic complexes:- CMI is bound to only one kind of ligands.



④ Heteroleptic complexes:- CMI is bound to more than one kind of ligands.
Ex $\rightarrow [Pt(NH_3)_2Cl_2]$

⑤ Coordination polyhedron:- Dabbe ki geometry (decided by CN of complex)

C.N

2 \rightarrow Linear

3 \rightarrow T planar

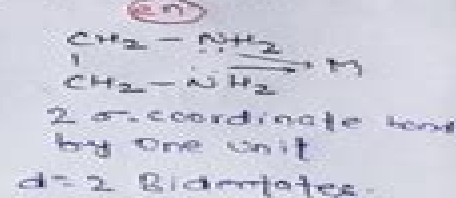
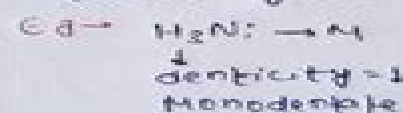
4 \rightarrow Sq planar

* 4 \rightarrow Tetrahedral

5 \rightarrow TBP

* 6 \rightarrow S.B.P or octahedral

Denticity:- no. of σ -coordinate bond formed by one unit of ligand.



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OAR Quarterly Index of Current Research Results United States. Air Force. Office of Aerospace Research,1967

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Bulletin mathématique Societatea de Științe Matematică,1988 **Russian Journal of Coordination Chemistry**

,1994-07 OAR Index of Research Results United States. Air Force. Office of Aerospace Research,1967 *The Publishers Weekly* ,1972 The Journal of Physical Chemistry ,1932 Reader's Guide to Periodical Literature Supplement ,1921

ISIS Cumulative Bibliography 1966-1975: Personalities and institutions History of Science Society,1980

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