

Name _____ Hour _____

Practice 17: Classifying and Balancing Chemical Reactions

- Balance each reaction (IF THEY ARE NOT BALANCED ALREADY).
- Classify each reaction as decomposition, synthesis, single replacement, double replacement, or combustion.

BALANCE REACTION	CLASSIFY REACTION
1. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$	_____
2. $\text{H}_2\text{SO}_4 + \text{Fe} \rightarrow \text{H}_2 + \text{FeSO}_4$	_____
3. $\text{Zn} + \text{S} \rightarrow \text{ZnS}$	_____
4. $\text{Al}_2\text{S}_3 \rightarrow \text{Al} + \text{S}$	_____
5. $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
6. $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu(OH)}_2$	_____
7. $\text{Cl}_2 + \text{KBr} \rightarrow \text{KCl} + \text{Br}_2$	_____
8. $\text{C}_4\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
9. $\text{Mg}_3(\text{PO}_4)_2 + \text{H}_2 \rightarrow \text{Mg} + \text{H}_3\text{PO}_4$	_____
10. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$	_____

Classifying Chemical Reactions Key Physical Science If8767

R. J. Stanley



Classifying Chemical Reactions Key Physical Science If8767:

Real World Examples of Four Types of Chemical Reactions , James W Smith developed Real World Examples of Four Types of Chemical Reactions a physical science and chemistry lesson for 9th grade students The students use the Internet to locate information about industrial pharmaceutical and environmental operations that use examples of the four general classes of chemical reactions These reactions are synthesis decomposition single replacement reactions and double replacement reactions The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K 12 classroom *Chemical Reactions and Their Equations* Ingo Waldemar Dagobert Hackh,1921

Chemical Reactions in Solvents and Melts G. Charlot,B. Trémillon,2013-10-22 Chemical Reactions in Solvents and Melts discusses the use of organic and inorganic compounds as well as of melts as solvents This book examines the applications in organic and inorganic chemistry as well as in electrochemistry Organized into two parts encompassing 15 chapters this book begins with an overview of the general properties and the different types of reactions including acid base reactions complex formation reactions and oxidation reduction reactions This text then describes the properties of inert and active solvents Other chapters consider the proton transfer reactions in polar solvents as well as the transfer of other ions This book discusses as well the solubility in a number of solvents by the formation of different bonds between the solute and the solvent molecule The final chapter deals with the general characteristics of the oxidation reduction reactions of melts This book is a valuable resource for chemists students and researchers *Gas-Phase Reactions* V.N. Kondratiev,E.E. Nikitin,2012-12-06 The present monograph appears after the death of Professor V N Kondratiev one of those scientists who have greatly contributed to the foundation of contemporary gas kinetics The most fundamental idea of chemical kinetics put forward at the beginning of the twentieth century and connected with names such as W Nernst M Bodenstein N N Semenov and C N Hinshelwood was that the complex chemical reactions are in fact a manifestation of a set of simpler elementary reactions involving but a small number of species V N Kondratiev was one of the first to adopt this idea and to start investigations on the elementary chemical reactions proper These investigations revealed explicitly that every elementary reaction in turn consisted of many elementary events usually referred to as elementary processes It took some time to realize that an elementary reaction represented in a very simple way by a macroscopic kinetic equation can be described on a microscopic level by a generalized Boltzmann equation Nevertheless up to the middle of the twentieth century gas kinetics was mainly concerned with the interpretation of complex chemical reactions via a set of elementary reactions But later on the situation changed drastically First the conditions for reducing microscopic equations to macroscopic ones were clearly set up These are essentially based on the fact that the small perturbations of the Maxwell Boltzmann distribution are caused by the reaction proper *Chemical Reactions* Gennadiï Efremovich Zaikov,Alfonso Jiménez,2004 Chemical Reactions Quantitative Level of Liquid Solid Phase Chemical Reactions in Gas, Liquid, and Solid Phases Gennadiï Efremovich

Zaikov, Ryszard Kozłowski, 2010 This book examines topical data on the subject of chemical reactions in different phases of gas liquid and solid states such as the classification of polymers in reactivity toward nitrogen oxide influence of the initiation rate of radicals on the kinetic characteristics of quercetin and dihydroquercetin in the methyl oleate oxidation supercritical carbon dioxide swelling of polyheteroarylenes synthesised in N-methylpyrrolidone inhibition of 2-hexenal oxidation by essential oils of ginger marjoram juniper berry black and white pepper specific properties of some biological composite materials properties and applications of aminoxyl radicals in polymer chemistry and others Quantitative Level of Chemical Reactions

Gennadii Efremovich Zaikov, Alfonso Jiménez, 2003 Using recent kinetic methods of investigations as well as such spectroscopic methods as electron spin resonance nuclear magnetic resonance and chemiluminescence 11 papers examine quantum level calculations of chemical reactions Presented by Zaikov Institute of Biochemical Physics Russia and Jimenez University of Alicante Spain the papers discuss such specific topics as the behaviour of molecules of low molecular substances in the presence of chain molecules ozone degradation of polyolefins with the example of paraffines and polyolefins interactions of surfactants and cellulose derivatives in homogenous and heterogeneous systems and transcrystallinity in natural fillers polypropylene composites **Chemical Reactions in Liquid and Solid Phase**

Gennadii Efremovich Zaikov, Alfonso Jiménez, 2003 This collection examines the influence of liquid and solid states during radical ionic and molecular reactions specifically how cage effect diffusion hindrance donor acceptor interaction electrostatic interaction dispersion forces and other factors affect the rates mechanism and direction of chemical reactions Topics of the 11 papers include spherical hydrogel particles for endovascular embolisation the synthesis and thermal degradation of phenolic resins mono molecular chain termination in dimethacrylate postpolymerisation and calculating the effect of chain deformation on macromolecule scission Most of the researchers work in Russia **Structure and Change** David M. Knight, 1983-01-01

Chemical Reaction Networks Oleg N. Temkin, Andrew V. Zeigarnik, D. G. Bonchev, 2019-12-02 Over the last decade increased attention to reaction dynamics combined with the intensive application of computers in chemical studies mathematical modeling of chemical processes and mechanistic studies has brought graph theory to the forefront of research It offers an advanced and powerful formalism for the description of chemical reactions and their intrinsic reaction mechanisms Chemical Reaction Networks A Graph Theoretical Approach elegantly reviews and expands upon graph theory as applied to mechanistic theory chemical kinetics and catalysis The authors explore various graph theoretical approaches to canonical representation numbering and coding of elementary steps and chemical reaction mechanisms the analysis of their topological structure the complexity estimation and classification of reaction mechanisms They discuss topologically distinctive features of multiroute catalytic and noncatalytic and chain reactions involving metal complexes With its careful balance of clear language and mathematical rigor the presentation of the authors significant original work and emphasis on practical applications and examples Chemical Reaction Networks A Graph Theoretical Approach is both an outstanding

reference and valuable tool for chemical research Collision Theory and Statistical Theory of Chemical Reactions S. G. Christov, Stefan G. Christov, 1980-07 Since the discovery of quantum mechanics more than fifty years ago the theory of chemical reactivity has taken the first steps of its development The knowledge of the electronic structure and the properties of atoms and molecules is the basis for an understanding of their interactions in the elementary act of any chemical process The increasing information in this field during the last decades has stimulated the elaboration of the methods for evaluating the potential energy of the reacting systems as well as the creation of new methods for calculation of reaction probabilities or cross sections and rate constants An exact solution to these fundamental problems of theoretical chemistry based on quantum mechanics and statistical physics however is still impossible even for the simplest chemical reactions Therefore different approximations have to be used in order to simplify one or the other side of the problem At present the basic approach in the theory of chemical reactivity consists in separating the motions of electrons and nuclei by making use of the Born Oppenheimer adiabatic approximation to obtain electronic energy as an effective potential for nuclear motion If the potential energy surface is known one can calculate in principle the reaction probability for any given initial state of the system The reaction rate is then obtained as an average of the reaction probabilities over all possible initial states of the reacting particles In the different stages of this calculational scheme additional approximations are usually introduced

CHEMICAL REACTIONS IN SOLVENTS AND MELTS Gaston Charlot, Bernard Trémillon, 1963 *The Reaction Path in Chemistry: Current Approaches and Perspectives* D. Heidrich, 2010-12-09 The so called reaction path RP with respect to the potential energy or the Gibbs energy free enthalpy is one of the most fundamental concepts in chemistry It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics It is particularly useful for the interpretation of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept *Chemical Reactions in Solvents and Melts* [by] G. Charlot and B. Trémillon. Translated by P. T. T. Harvey Gaston Charlot, B. Trémillon, 1969

Motion and Energy, Chemical Reactions Karl Dittmer, Michael Fiasca, Portland Project Committee, 1970 **Rates of**

chemical reactions R. J. Stanley, 1971

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