

Name _____ Hour _____

Practice 17: Classifying and Balancing Chemical Reactions

- Balance each reaction (IF THEY ARE NOT BALANCED ALREADY).
- Classify each reaction as decomposition, synthesis, single replacement, double replacement, or combustion.

BALANCE REACTION	CLASSIFY REACTION
1. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$	_____
2. $\text{H}_2\text{SO}_4 + \text{Fe} \rightarrow \text{H}_2 + \text{FeSO}_4$	_____
3. $\text{Zn} + \text{S} \rightarrow \text{ZnS}$	_____
4. $\text{Al}_2\text{S}_3 \rightarrow \text{Al} + \text{S}$	_____
5. $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
6. $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu(OH)}_2$	_____
7. $\text{Cl}_2 + \text{KBr} \rightarrow \text{KCl} + \text{Br}_2$	_____
8. $\text{C}_4\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
9. $\text{Mg}_3(\text{PO}_4)_2 + \text{H}_2 \rightarrow \text{Mg} + \text{H}_3\text{PO}_4$	_____
10. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$	_____

Classifying Chemical Reaction Physical Science If8767

T. E. Graedel



Classifying Chemical Reaction Physical Science If8767:

Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh, 1921 *Real World Examples of Four Types of Chemical Reactions*, James W Smith developed Real World Examples of Four Types of Chemical Reactions a physical science and chemistry lesson for 9th grade students The students use the Internet to locate information about industrial pharmaceutical and environmental operations that use examples of the four general classes of chemical reactions These reactions are synthesis decomposition single replacement reactions and double replacement reactions The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K 12 classroom

Gas-Phase Reactions V.N. Kondratiev, E.E. Nikitin, 2012-12-06 The present monograph appears after the death of Professor V N Kondratiev one of those scientists who have greatly contributed to the foundation of contemporary gas kinetics The most fundamental idea of chemical kinetics put forward at the beginning of the twentieth century and connected with names such as W Nernst M Bodenstein N N Semenov and C N Hinshelwood was that the complex chemical reactions are in fact a manifestation of a set of simpler elementary reactions involving but a small number of species V N Kondratiev was one of the first to adopt this idea and to start investigations on the elementary chemical reactions proper These investigations revealed explicitly that every elementary reaction in turn consisted of many elementary events usually referred to as elementary processes It took some time to realize that an elementary reaction represented in a very simple way by a macroscopic kinetic equation can be described on a microscopic level by a generalized Boltzmann equation Nevertheless up to the middle of the twentieth century gas kinetics was mainly concerned with the interpretation of complex chemical reactions via a set of elementary reactions But later on the situation changed drastically First the conditions for reducing microscopic equations to macroscopic ones were clearly set up These are essentially based on the fact that the small perturbations of the Maxwell Boltzmann distribution are caused by the reaction proper *Collision Theory and Statistical Theory of Chemical Reactions* S. G. Christov, Stefan G. Christov, 1980-07 Since the discovery of quantum mechanics more than fifty years ago the theory of chemical reactivity has taken the first steps of its development The knowledge of the electronic structure and the properties of atoms and molecules is the basis for an understanding of their interactions in the elementary act of any chemical process The increasing information in this field during the last decades has stimulated the elaboration of the methods for evaluating the potential energy of the reacting systems as well as the creation of new methods for calculation of reaction probabilities or cross sections and rate constants An exact solution to these fundamental problems of theoretical chemistry based on quantum mechanics and statistical physics however is still impossible even for the simplest chemical reactions Therefore different approximations have to be used in order to simplify one or the other side of the problem At present the basic approach in the theory of chemical reactivity consists in separating the motions of electrons and nuclei by making use of the Born Oppenheimer adiabatic approximation to obtain electronic energy as an effective potential

for nuclear motion If the potential energy surface is known one can calculate in principle the reaction probability for any given initial state of the system The reaction rate is then obtained as an average of the reaction probabilities over all possible initial states of the reacting articles In the different stages of this calculational scheme additional approximations are usually introduced

Chemical Reactivity in Liquids Michael Moreau, Pierre Turq, 2012-12-06 Understanding chemical reactivity has been the permanent concern of chemists from time immemorial If we were able to understand it and express it quantitatively there would practically remain no unsolved mystery and reactions would be fully predictable with their products and rates and even side reactions The beautiful developments of thermodynamics through the 19th century supplied us with the knowledge of the way a reactions progresses and the statistical view initiated by Gibbs has progressively led to an understanding closer to the microscopic phenomena But it was always evident to all that these advances still left our understanding of chemical reactivity far behind our empirical knowledge of the chemical reaction in its practically infinite variety The advances of recent years in quantum chemistry and statistical mechanics enhanced by the present availability of powerful and fast computers are very fast changing this picture and bringing us really close to a microscopic understanding of chemical equilibria reaction rates etc This is the reason why our Society encouraged a few years ago the initiative of Professor Savo Bratos who with a group of French colleagues prepared an impressive study on *Reactivite chimique en phase liquide* a prospective report which was jointly published by the Societe Fran

THEORY CHEM REACT DYN Michael Baer, 1985-11-19 The Reaction Path in Chemistry: Current Approaches and Perspectives D.

Heidrich, 2010-12-09 The so called reaction path RP with respect to the potential energy or the Gibbs energy free enthalpy is one of the most fundamental concepts in chemistry It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics It is particularly useful for the interpretation of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept

Theory of Chemical Reaction Dynamics Antonio Laganà, 2005

Oscillations in Chemical Reactions, 1983-10-01

Reaction In Condensed Phases Henry Eyring, 2012-12-02 Physical

Chemistry An Advanced Treatise Reactions in Condensed Phases Volume VII deals with reactions in condensed phases The purpose of this treatise is to present a comprehensive treatment of physical chemistry for advanced students and investigators in a reasonably small number of volumes An attempt has been made to include all important topics in physical chemistry together with borderline subjects which are of particular interest and importance The book begins by discussing the basic principles of reaction rates in solution This is followed by separate chapters on estimating the rate parameters of elementary reactions the use of correlation diagrams to interpret organic reactions perturbation of reaction rates by substituents and inorganic reactions Subsequent chapters cover the important field of free radicals including chain reactions and solvent effects heterogeneous catalysis various types of surface reactions surface annealing electron reactions nucleation and radiation chemistry The book presents a broad picture of current developments in reaction rates in condensed phases in a form accessible to all students of chemical kinetics This treatment by experts in widely different areas will hopefully meet many student needs and provide a useful overview for all

THEORY CHEM REACT DYN Michael Baer, 1985-11-19

Oscillations in Chemical Reactions, 2014-04-17 *Collisions and Chemical Reactions of Gases* Open University, Open University. ST28- Course Team, 1973 **Chemical Reactions** Gennadii Efremovich Zaikov, Alfonso Jiménez, 2004 Chemical Reactions Quantitative Level of Liquid Solid Phase *Structure and Change* David M. Knight, 1983-01-01 Chemical Reactions in Solvents and Melts G. Charlot, B. Trémillon, 2013-10-22 Chemical Reactions in Solvents and Melts discusses the use of organic and inorganic compounds as well as of melts as solvents This book examines the applications in organic and inorganic chemistry as well as in electrochemistry Organized into two parts encompassing 15 chapters this book begins with an overview of the general properties and the different types of reactions including acid base reactions complex formation reactions and oxidation reduction reactions This text then describes the properties of inert and active solvents Other chapters consider the proton transfer reactions in polar solvents as well as the transfer of other ions This book discusses as well the solubility in a number of solvents by the formation of different bonds between the solute and the solvent molecule The final chapter deals with the general characteristics of the oxidation reduction reactions of melts This book is a valuable resource for chemists students and researchers

Chemical Compounds in the Atmosphere T. E. Graedel, 1978 Chemical Compounds in The Atmosphere **Chemical Reactions in Complex Mixtures** Ajit V. Sapre, Frederick J. Krambeck, 1991-07-25 In recent years there has been a convergence of trends in chemical reaction engineering and chemistry which have set the stage for significant advances in kinetic and thermodynamic modeling of processes New analytical chemistry methods new mathematical methods and new computational tools facilitate a more fundamental approach and a deeper understanding of chemical reactions in complex mixtures with very large numbers of compounds such as petroleum fractions This fortunate state of affairs has stimulated important new work both in academia and industrial research labs The purpose of the workshop that led to this book was to bring together researchers at the forefront of this field

to review the state of the art stimulate communication and cooperation between industry and academia and develop a cohesive picture of research trends and future directions The chapters of the book have been organized into four main areas Continuous mixtures where the very large numbers of discrete compounds present are regarded as making up a continuum Structure activity relationships where the nature and rates of the reactions that a particular molecule undergoes are correlated with its chemical structure thus allowing the kinetics of very large numbers of compounds to be described by a few parameters Kinetic analysis where mathematical techniques are applied to analyze the behavior of kinetic networks and Thermodynamics emphasizing the practical and computational aspects of chemical equilibrium in complex mixtures

The Theory of Chemical Reaction Dynamics D. C. Clary, 1986-03-31 **Chemical Reactions** Kaufman George Falk, 1920

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