

Name _____ Hour _____

Practice 17: Classifying and Balancing Chemical Reactions

- Balance each reaction (IF THEY ARE NOT BALANCED ALREADY).
- Classify each reaction as decomposition, synthesis, single replacement, double replacement, or combustion.

BALANCE REACTION	CLASSIFY REACTION
1. $\text{NH}_3 + \text{HCl} \rightarrow \text{NH}_4\text{Cl}$	_____
2. $\text{H}_2\text{SO}_4 + \text{Fe} \rightarrow \text{H}_2 + \text{FeSO}_4$	_____
3. $\text{Zn} + \text{S} \rightarrow \text{ZnS}$	_____
4. $\text{Al}_2\text{S}_3 \rightarrow \text{Al} + \text{S}$	_____
5. $\text{C}_{12}\text{H}_{22}\text{O}_{11} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
6. $\text{NaOH} + \text{CuSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{Cu(OH)}_2$	_____
7. $\text{Cl}_2 + \text{KBr} \rightarrow \text{KCl} + \text{Br}_2$	_____
8. $\text{C}_4\text{H}_{12} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$	_____
9. $\text{Mg}_3(\text{PO}_4)_2 + \text{H}_2 \rightarrow \text{Mg} + \text{H}_3\text{PO}_4$	_____
10. $\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2\text{O}$	_____

Classifying Chemical Reactions Physical Science If8767

**Carl Alfred Jacobson, Clifford A.
Hampel**



Classifying Chemical Reactions Physical Science If8767:

Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh, 1921 **Real World Examples of Four Types of Chemical Reactions**, James W Smith developed Real World Examples of Four Types of Chemical Reactions a physical science and chemistry lesson for 9th grade students The students use the Internet to locate information about industrial pharmaceutical and environmental operations that use examples of the four general classes of chemical reactions These reactions are synthesis decomposition single replacement reactions and double replacement reactions The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K 12 classroom

Chemical Reactivity in Liquids Michael Moreau, Pierre Turq, 2012-12-06 Understanding chemical reactivity has been the permanent concern of chemists from time immemorial If we were able to understand it and express it quantitatively there would practically remain no unsolved mystery and reactions would be fully predictable with their products and rates and even side reactions The beautiful developments of thermodynamics through the 19th century supplied us with the knowledge of the way a reactions progresses and the statistical view initiated by Gibbs has progressively led to an understanding closer to the microscopic phenomena But it was always evident to all that these advances still left our understanding of chemical reactivity far behind our empirical knowledge of the chemical reaction in its practically infinite variety The advances of recent years in quantum chemistry and statistical mechanics enhanced by the present availability of powerful and fast computers are very fast changing this picture and bringing us really close to a microscopic understanding of chemical equilibria reaction rates etc This is the reason why our Society encouraged a few years ago the initiative of Professor Savo Bratos who with a group of French colleagues prepared an impressive study on *Reactivite chimique en phase liquide* a prospective report which was jointly published by the Societe Fran Gas-Phase Reactions V.N. Kondratiev, E.E. Nikitin, 2012-12-06 The present monograph appears after the death of Professor V N Kondratiev one of those scientists who have greatly contributed to the foundation of contemporary gas kinetics The most fundamental idea of chemical kinetics put forward at the beginning of the twentieth century and connected with names such as W Nernst M Bodenstein N N Semenov and C N Hinshelwood was that the complex chemical reactions are in fact a manifestation of a set of simpler elementary reactions involving but a small number of species V N Kondratiev was one of the first to adopt this idea and to start investigations on the elementary chemical reactions proper These investigations revealed explicitly that every elementary reaction in turn consisted of many elementary events usually referred to as elementary processes It took some time to realize that an elementary reaction represented in a very simple way by a macroscopic kinetic equation can be described on a microscopic level by a generalized Boltzmann equation Nevertheless up to the middle of the twentieth century gas kinetics was mainly concerned with the interpretation of complex chemical reactions via a set of elementary reactions But later on the situation changed drastically First the conditions for reducing microscopic equations to macroscopic ones were clearly set up These are essentially based

on the fact that the small perturbations of the Maxwell Boltzmann distribution are caused by the reaction proper

THEORY CHEM REACT DYN Michael Baer,1985-11-19

Collision Theory and Statistical Theory of Chemical

Reactions S. G. Christov,Stefan G. Christov,1980-07 Since the discovery of quantum mechanics more than fifty years ago the theory of chemical reactivity has taken the first steps of its development The knowledge of the electronic structure and the properties of atoms and molecules is the basis for an understanding of their interactions in the elementary act of any chemical process The increasing information in this field during the last decades has stimulated the elaboration of the methods for evaluating the potential energy of the reacting systems as well as the creation of new methods for calculation of reaction probabilities or cross sections and rate constants An exact solution to these fundamental problems of theoretical chemistry based on quantum mechanics and statistical physics however is still impossible even for the simplest chemical reactions Therefore different approximations have to be used in order to simplify one or the other side of the problem At present the basic approach in the theory of chemical reactivity consists in separating the motions of electrons and nuclei by making use of the Born Oppenheimer adiabatic approximation to obtain electronic energy as an effective potential for nuclear motion If the potential energy surface is known one can calculate in principle the reaction probability for any given initial state of the system The reaction rate is then obtained as an average of the reaction probabilities over all possible initial states of the reacting particles In the different stages of this calculational scheme additional approximations are usually introduced

THEORY CHEM REACT DYN Michael Baer,1985-11-19

Theory of Chemical Reaction Dynamics

Antonio Laganà,2005

The Reaction Path in Chemistry: Current Approaches and Perspectives D. Heidrich,2010-12-09

The so called reaction path RP with respect to the potential energy or the Gibbs energy free enthalpy is one of the most fundamental concepts in chemistry It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction This concept is an implicit component of conventional transition state theory TST The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination However an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics It is particularly useful for the interpretation of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept

Encyclopedia of Chemical Reactions Carl Alfred Jacobson,Clifford A. Hampel,1956

Chemical

Reactions Kaufman George Falk,1920 *Chemical Reactions in Solvents and Melts* G. Charlot,B. Trémillon,2013-10-22
 Chemical Reactions in Solvents and Melts discusses the use of organic and inorganic compounds as well as of melts as solvents This book examines the applications in organic and inorganic chemistry as well as in electrochemistry Organized into two parts encompassing 15 chapters this book begins with an overview of the general properties and the different types of reactions including acid base reactions complex formation reactions and oxidation reduction reactions This text then describes the properties of inert and active solvents Other chapters consider the proton transfer reactions in polar solvents as well as the transfer of other ions This book discusses as well the solubility in a number of solvents by the formation of different bonds between the solute and the solvent molecule The final chapter deals with the general characteristics of the oxidation reduction reactions of melts This book is a valuable resource for chemists students and researchers *Structure and Change* David M. Knight,1983-01-01 **Chemical Reactions** Gennadii Efremovich Zaikov,Alfonso Jiménez,2004
 Chemical Reactions Quantitative Level of Liquid Solid Phase Chemical Reactions in Liquid and Solid Phase Gennadii Efremovich Zaikov,Alfonso Jiménez,2003 This collection examines the influence of liquid and solid states during radical ionic and molecular reactions specifically how cage effect diffusion hindrance donor acceptor interaction electrostatic interaction dispersion forces and other factors affect the rates mechanism and direction of chemical reactions Topics of the 11 papers include spherical hydrogel particles for endovascular embolisation the synthesis and thermal degradation of phenolic resins mono molecular chain termination in dimethacrylate postpolymerisation and calculating the effect of chain deformation on macromolecule scission Most of the researchers work in Russia **Physical Chemistry of Fast Reactions** Bryan Philip Levitt,1973 **Chemical Reactions in Gas, Liquid, and Solid Phases** Gennadii Efremovich Zaikov,Ryszard Kozłowski,2010 The Theory of Chemical Reaction Dynamics D. C. Clary,1986-03-31 Chemical Energetics and the Curriculum Douglas James Millen,1969 *Chemical Reactions* ,1995

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