

## Chapter 1

### Some Basic Concepts of Chemistry

#### Question 1.1:

Calculate the molar mass of the following:

- (i)  $\text{H}_2\text{O}$                       (ii)  $\text{CO}_2$                       (iii)  $\text{CH}_4$

#### Answer 1.1:

- (i)  $\text{H}_2\text{O}$ :

The molecular mass of water,  $\text{H}_2\text{O}$   
 $= (2 \times \text{Atomic mass of hydrogen}) + (1 \times \text{Atomic mass of oxygen})$   
 $= [2(1.0084) + 1(16.00 \text{ u})]$   
 $= 2.016 \text{ u} + 16.00 \text{ u} = 18.016 = 18.02 \text{ u}$

- (ii)  $\text{CO}_2$ :

The molecular mass of carbon dioxide,  $\text{CO}_2$   
 $= (1 \times \text{Atomic mass of carbon}) + (2 \times \text{Atomic mass of oxygen})$   
 $= [1(12.011 \text{ u}) + 2(16.00 \text{ u})]$   
 $= 12.011 \text{ u} + 32.00 \text{ u} = 44.01 \text{ u}$

- (iii)  $\text{CH}_4$ :

The molecular mass of methane,  $\text{CH}_4$   
 $= (1 \times \text{Atomic mass of carbon}) + (4 \times \text{Atomic mass of hydrogen})$   
 $= [1(12.011 \text{ u}) + 4(1.008 \text{ u})]$   
 $= 12.011 \text{ u} + 4.032 \text{ u} = 16.043 \text{ u}$

#### Question 1.2:

Calculate the mass per cent of different elements present in sodium sulphate ( $\text{Na}_2\text{SO}_4$ ).

#### Answer 1.2:

The molecular formula of sodium sulphate is ( $\text{Na}_2\text{SO}_4$ ).

Molar mass of  $\text{Na}_2\text{SO}_4 = [(2 \times 23.0) + (32.066) + 4(16.00)] = 142.066 \text{ g}$

Mass percent of an element =  $\frac{\text{Mass of that element in the compound}}{\text{Molar Mass of the compound}} \times 100$

$\therefore$  Mass percent of sodium:  $\frac{46.0 \text{ g}}{142.066 \text{ g}} \times 100 = 32.379 = 32.4\%$

Mass percent of sulphur:  $\frac{32.066 \text{ g}}{142.066 \text{ g}} \times 100 = 22.57 = 22.6\%$

Mass percent of oxygen:  $\frac{64.0 \text{ g}}{142.066 \text{ g}} \times 100 = 45.049 = 45.05\%$

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**Lalit Goel**



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### **Table of Contents Class 11 Chemistry Ncert Solutions Chapter**

1. Understanding the eBook Class 11 Chemistry Ncert Solutions Chapter
  - The Rise of Digital Reading Class 11 Chemistry Ncert Solutions Chapter
  - Advantages of eBooks Over Traditional Books
2. Identifying Class 11 Chemistry Ncert Solutions Chapter
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Class 11 Chemistry Ncert Solutions Chapter
  - User-Friendly Interface
4. Exploring eBook Recommendations from Class 11 Chemistry Ncert Solutions Chapter
  - Personalized Recommendations

- 
- Class 11 Chemistry Ncert Solutions Chapter User Reviews and Ratings
  - Class 11 Chemistry Ncert Solutions Chapter and Bestseller Lists
5. Accessing Class 11 Chemistry Ncert Solutions Chapter Free and Paid eBooks
    - Class 11 Chemistry Ncert Solutions Chapter Public Domain eBooks
    - Class 11 Chemistry Ncert Solutions Chapter eBook Subscription Services
    - Class 11 Chemistry Ncert Solutions Chapter Budget-Friendly Options
  6. Navigating Class 11 Chemistry Ncert Solutions Chapter eBook Formats
    - ePub, PDF, MOBI, and More
    - Class 11 Chemistry Ncert Solutions Chapter Compatibility with Devices
    - Class 11 Chemistry Ncert Solutions Chapter Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Class 11 Chemistry Ncert Solutions Chapter
    - Highlighting and Note-Taking Class 11 Chemistry Ncert Solutions Chapter
    - Interactive Elements Class 11 Chemistry Ncert Solutions Chapter
  8. Staying Engaged with Class 11 Chemistry Ncert Solutions Chapter
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Class 11 Chemistry Ncert Solutions Chapter
  9. Balancing eBooks and Physical Books Class 11 Chemistry Ncert Solutions Chapter
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Class 11 Chemistry Ncert Solutions Chapter
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Class 11 Chemistry Ncert Solutions Chapter
    - Setting Reading Goals Class 11 Chemistry Ncert Solutions Chapter
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Class 11 Chemistry Ncert Solutions Chapter
    - Fact-Checking eBook Content of Class 11 Chemistry Ncert Solutions Chapter

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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