Question 1.26:
If the density of some lake water is 1.25 g mL^{-1} and contains 92 g of Na^+ ions per kg of water, calculate the molality of Na^+ ions in the lake.

Answer 1.26:
Number of moles present in 92 g of Na^+ ions = $\frac{92\ g}{23\ g\ mol^{-1}} = 4\ mol$

Therefore, molality of Na^+ ions in the lake = $\frac{4\ mol}{1\ kg} = 4\ m$

Question 1.27:
If the solubility product of CuS is 6×10^{-16}, calculate the maximum molarity of CuS in aqueous solution.

Answer 1.27:
Solubility product of CuS, $K_{sp} = 6 \times 10^{-16}$ Let s be the solubility of CuS in mol L^{-1}.

$$CuS \leftrightarrow \underset{s}{C^{2+}} + \underset{s}{S^{2+}} + S^{2-}$$

$$K_{sp} = [Cu^{2+}][S^{2-}]$$

Now,
$= s \times s = s^2$
Then, we have, $K_{sp} = s^2 = 6 \times 10^{-16}$
$\Rightarrow s = \sqrt{6 \times 10^{-16}} = 2.45 \times 10^{-8}$ mol L^{-1}
Hence, the maximum molarity of CuS in an aqueous solution is 2.45×10^{-8} mol L^{-1}.

Question 1.28:
Calculate the mass percentage of aspirin ($C_9H_8O_4$) in acetonitrile (CH_3CN) when 6.5 g of $C_9H_8O_4$ is dissolved in 450 g of CH_3CN.

Answer 1.28:
6.5 g of $C_9H_8O_4$ is dissolved in 450 g of CH_3CN.
Then, total mass of the solution = (6.5 + 450) g = 456.5 g

Therefore, mass percentage of $C_9H_8O_4$ = $\frac{6.5}{456.5} \times 100\ \% = 1.424\%$

Question 1.29:
Nalorphene ($C_{19}H_{21}NO_3$), similar to morphine, is used to combat withdrawal symptoms in narcotic users. Dose of nalorphene generally given is 1.5 mg. Calculate the mass of 1.5×10^{-3}m aqueous solution required for the above dose.

Answer 1.29:
The molar mass of Nalorphene ($C_{19}H_{21}NO_3$) is given as:
$19 \times 12 + 12 \times 1 + 1 \times 14 + 3 \times 16 = 311\ g\ mol^{-1}$
In 1.5×10^{-3}m aqueous solution of nalorphene, 1 kg (1000 g) of water contains 1.5×10^{-3} mol
$= 1.5 \times 10^{-3} \times 311\ g = 0.4665\ g$
Therefore, total mass of the solution = (1000 + 0.4665) g = 1000.4665 g
This implies that the mass of the solution containing 0.4665 g of nalorphene is 1000.4665 g.
Therefore, mass of the solution containing 1.5 mg of nalorphene is: $\frac{1000.4665 \times 1.5 \times 10^{-3}}{0.4665}\ g = 3.22\ g$
Hence, the mass of aqueous solution required is 3.22 g.

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Table of Contents Class 1chemistry Ncert Solutions Chapter

1. Understanding the eBook Class 1chemistry Ncert Solutions Chapter
    - The Rise of Digital Reading Class 1chemistry Ncert Solutions Chapter
    - Advantages of eBooks Over Traditional Books
2. Identifying Class 1chemistry Ncert Solutions Chapter
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Class 1chemistry Ncert Solutions Chapter
    - User-Friendly Interface
4. Exploring eBook Recommendations from Class 1chemistry Ncert Solutions Chapter
    - Personalized Recommendations
    - Class 1chemistry Ncert Solutions Chapter User Reviews and Ratings
    - Class 1chemistry Ncert Solutions Chapter and Bestseller Lists
5. Accessing Class 1chemistry Ncert Solutions Chapter Free and Paid eBooks
    - Class 1chemistry Ncert Solutions Chapter Public Domain eBooks
    - Class 1chemistry Ncert Solutions Chapter eBook Subscription Services
    - Class 1chemistry Ncert Solutions Chapter Budget-Friendly Options

6. Navigating Class 1chemistry Ncert Solutions Chapter eBook Formats
    - ePub, PDF, MOBI, and More
    - Class 1chemistry Ncert Solutions Chapter Compatibility with Devices
    - Class 1chemistry Ncert Solutions Chapter Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Class 1chemistry Ncert Solutions Chapter
    - Highlighting and Note-Taking Class 1chemistry Ncert Solutions Chapter
    - Interactive Elements Class 1chemistry Ncert Solutions Chapter
8. Staying Engaged with Class 1chemistry Ncert Solutions Chapter
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Class 1chemistry Ncert Solutions Chapter
9. Balancing eBooks and Physical Books Class 1chemistry Ncert Solutions Chapter
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Class 1chemistry Ncert Solutions Chapter
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Class 1chemistry Ncert Solutions Chapter
    - Setting Reading Goals Class 1chemistry Ncert Solutions Chapter
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Class 1chemistry Ncert Solutions Chapter
    - Fact-Checking eBook Content of Class 1chemistry Ncert Solutions Chapter
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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