

## CHAPTER

## 02

## SOLUTION

1. **Solution** : A homogeneous mixture of two or more substances is known as solution.
2. **Solute** : The substance present in smaller amount in a solution is called solute.
3. **Solvent** : The substance present in larger amount in a solution is called solvent.
4. **Binary solution** : Those solutions which contain two components are called binary solutions, e.g., salt solution, benzene and toluene.
5. **Ternary solutions** : Those solutions which contain three components are called ternary solutions, e.g., ethanol + water + acetic acid.
6. **Aqueous solution** : When solute is dissolved in water, it is called aqueous solution, e.g., sugar solution, ethanol in water.
7. **Non-aqueous solution** : When solute is dissolved in solvent other than water, it is called non-aqueous solution, e.g., iodine dissolved in alcohol (Tincture of iodine).
8. **Weight Fraction** =  $\frac{\text{weight of solute}}{\text{weight of solution}}$
9. **Weight percentage**  

$$= \frac{\text{weight of solute}}{\text{weight of solution}} \times 100$$

Weight of solute present in 100 grams of solution is called weight percentage

Volume of solute present in 100 ml of the solution is called volume percentage
10. **Volume percentage**  

$$= \frac{\text{Volume of solute}}{\text{Volume of solution}} \times 100$$
11. **Volume Fraction** =  $\frac{\text{Volume of solute}}{\text{Volume of solution}}$
12. **Molarity (M)**  

$$= \frac{\text{No of moles of solute}}{\text{Volume of solution in liters}}$$

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1}{V(\text{lit})}$$
13. **Molarity** =  $\frac{\text{No of moles of solute}}{\text{Volume of solution in ml} \times \frac{1000}{1000}}$
14. **Molarity** =  $\frac{\text{Weight of solute}}{\text{GMWt of solute}} \times \frac{1000}{V \text{ in ml}}$   

$$(\text{or}) M = \frac{W}{GMW} \times \frac{1000}{V(\text{ml})}$$
15. **Weight of solute** =  $\frac{M \times GMW \times V \text{ in ml}}{1000}$
16. **Molarity for dilute solutions** :  $M_1 V_1 = M_2 V_2$
17. **For neutralisation reactions** :  $\frac{M_1 V_1}{n_1} = \frac{M_2 V_2}{n_2}$
18. **For mixed solution molarity**  

$$= \frac{M_1 V_1 + M_2 V_2 + \dots}{V_1 + V_2 + \dots}$$

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