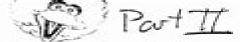
CLASSIFYING CHEMICAL REACTIONS

14. SO₃ + H₂O → H₂SO₄



Classify the following reactions as synthesis, decomposition, single replacement or double replacement.

repi	acement.	ANSW	EKS/
1.	2KCIO ₃ → 2KCI + 3O ₂	De con	~position
2.	HCI + NaOH → NaCI + H ₂ O	Double	Replacement
з.	Mg + 2HCl → MgCl ₂ + H ₂	Single	Replacement.
4.	2H ₂ + O ₂ → 2H ₂ O	Synth	esis
5.	2AI + 3NIBr₂ → 2AIBr₃ + 3NI	Single	Replacement
6.	4AI + 3O₂ → 2AI₂O₃	Synth	resis
7.	2NaCl → 2Na + Cl ₂	Decom	position
8.	$CaCl_2 + F_2 \rightarrow CaF_2 + Cl_2$	Single	Replacement
9.	AgNO ₃ + KCI → AgCI + KNO ₃	Double	Replacement
10.	N ₂ + 3H ₂ → 2NH ₃	Synth	nesis
11.	2H ₂ O ₂ → 2H ₂ O + O ₂	Decomp	positron
12.	$(NH_4)_2SO_4 + Ba(NO_3)_2 \rightarrow BaSO_4$	+ 2NH4NO, DO	uble Replacement
13.	Mgl ₂ + Br ₂ → MgBr ₂ + l ₂	Single	Replacement

Synthesis

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Lifeliqe

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Real World Examples of Four Types of Chemical Reactions , James W Smith developed Real World Examples of Four Types of Chemical Reactions a physical science and chemistry lesson for 9th grade students The students use the Internet to locate information about industrial pharmaceutical and environmental operations that use examples of the four general classes of chemical reactions These reactions are synthesis decomposition single replacement reactions and double replacement reactions The Louisiana Challenge Grant provides the lesson online as part of the Louisiana Challenge collection of activities for the K 12 classroom

Chemical Reactions and Their Equations Ingo Waldemar Dagobert Hackh,1921

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