

Chapter - 3

Classification Of Elements And Periodicity In

Properties

The first systematic classification of elements was provided by Russian chemist D.I. Mendeleev.

1. Mendeleev's periodic law

"The physical and chemical properties of elements are periodic functions of their atomic weight."

2. It was modified to Modern Periodic law:

"The physical and chemical properties of elements are periodic functions of their atomic numbers."

It is the long form of periodic table:

7 Horizontal rows are called Periods and 18 Vertical columns are called Group

Group-1 are called **Alkali metals**Group-15 are called **Pnicogens**Group-17 are called **Halogens**Group-18 are called **Noble gases**

- 3. 1st period—2 elements 2nd and 3rd period—8 elements 4th and 5th period—18 elements 6th period—32 elements 7th period—Incomplete (32 elements)
- 4. Groups

1 and 2 — 's' block elements last electron entered in 's' subshell [s¹, s²] 3 to 12 — 'd' block elements last electrons entered in 'd' subshell [d¹ to d¹⁰]. 13 to 18 — 'p' block elements last electrons enter in 'p' subshell [p¹ to p⁶]. Two f-block series lanthanoids and actinoids are placed in the bottom of periodic table.

5. (A) In 's' and 'p' block elements the electrons enters in outer most shell. In 'd' block elements the electron enters in the penultimate shell (n-1).

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