

CLEP chemistry LATEST 2025 UPDATE!!!

Which chemical equation represents this reaction:

a Bronsted-Lowry acid-base reaction

- A.) $2 \text{KClO}_3(\text{s}) \rightarrow 2 \text{KClO}(\text{s}) + 3 \text{O}_2(\text{g})$
- B.) $\text{Zn}(\text{s}) + \text{Cu}^{2+}(\text{aq}) \rightarrow \text{Zn}^{2+}(\text{aq}) + \text{Cu}(\text{s})$
- C.) $\text{CO}^{3+}(\text{aq}) + 6 \text{NH}_3(\text{aq}) \rightarrow [\text{Co}(\text{NH}_3)_6]^{3+}(\text{aq})$
- D.) $\text{H}_2\text{O}(\text{l}) + \text{NH}_3(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{NH}_3(\text{aq})$

The attractions between atoms within a molecule of NH_3 are best characterized as

- A. hydrogen bonds
- B. ionic bonds
- C. polar covalent bonds
- D. London (dispersion) forces
- E. ion-dipole forces - polar covalent bonds



The reaction mixture above is at equilibrium at 298 K. The value of the equilibrium constant for the reaction is 16 at 298K. If $\text{R}(\text{g})$, $\text{X}(\text{g})$ and $\text{Z}(\text{g})$ each have an equilibrium concentration of 2.0 M, what is the equilibrium concentration of $\text{T}(\text{g})$? - 0.50

7) what is the pH of a solution made by diluting 1.00 ml of 0.10 M HCL with enough distilled water to make 1.00L of solution? - 4.0

1) The electron configuration notation for molybdenum is

- A) $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^6$
- B) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 3f^6$
- C) $1s^2 2s^2 2p^6 2d^{10} 3s^2 3p^6 3d^{10} 3f^4$
- D) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^4$
- E) $1s^2 2s^2 3s^2 4s^2 5s^2 2p^6 3p^6 4p^6 3d^{10} 4d^4$ - D:) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^4$. The order of filling is $1s 2s 2p 3s 3p 4s 3d 4p 5s 4d 5p 6s 4f 5d 6p 7s 5f 6d 7p$ with s sublevels holding a maximum of 2 electrons, p holding up to 6, d up to 10 and f up to 14.

2) The periodic property which decreases all the way across the table is

- A) Ionization energy
- B) Atomic radius
- C) Ionic size

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SA Adler



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