

## Balancing Act

Name \_\_\_\_\_

{KEY}

Atoms are not created or destroyed during a chemical reaction. Scientists know that there must be the same number of atoms on each side of the reaction. To balance the chemical equation, you must add coefficients in front of the chemical formulas in the equation. You cannot add or change subscripts!

- 1) Determine number of atoms for each element.
- 2) Pick an element that is not equal on both sides of the equation.
- 3) Add a coefficient in front of the formula with that element and adjust your counts.
- 4) Continue adding coefficients to get the same number of atoms of each element on each side.



$$\text{Mg} = 2$$

$$\text{Mg} = 2$$

$$\text{O} = 2$$

$$\text{O} = 2$$

Try these:



$$\text{Ca} = 1 \quad 2$$

$$\text{Ca} = 1 \quad 2$$

$$\text{O} = 2 \quad 2$$

$$\text{O} = 1 \quad 2$$



$$\text{N} = 2 \quad 2$$

$$\text{N} = 1 \quad 2$$

$$\text{H} = 2 \quad 6$$

$$\text{H} = 3 \quad 6$$



$$\text{Cu} = 2 \quad 4$$

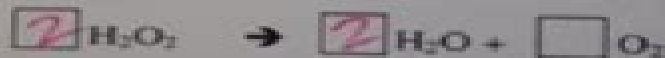
$$\text{Cu} = 1 \quad 4$$

$$\text{O} = 1 \quad 2$$

$$\text{O} = 2 \quad 2$$

$$\text{C} = 1 \quad 1$$

$$\text{C} = 1 \quad 1$$



$$\text{H} = 2 \quad 4$$

$$\text{H} = 2 \quad 4$$

$$\text{O} = 2 \quad 4$$

$$\text{O} = 3 \quad 4$$

# Balancing Act Practice Science Spot Answers

**Brendan G. Carr**



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