## **Balancing Chemical Equations**

Once we have a written a chemical equation, we must make sure to balance it! Balancing a chemical equation means to make sure that there are the same number of each type of atom on both sides of the chemical equation. So, if there are five carbons on the left side (reactants), there needs to be five carbons on the right side (products).

So, how do we balance an equation? Follow the following steps...

- a. Change only stoichiometric coefficients those big numbers in front of the compounds (not subscripts in the formulas – that would change the identity of the compound!)
- b. Start with most complex
- If found in more than one place on each side, do it last
- d. Watch out for odd numbers (make them into even ones if needed)
- e. Check your work!

Now, try to balance the following equations. In addition, name the bolded compounds for some nomenclature practice!

1.  $NH_4NO_3(s) \rightarrow N_2O(g) + 2H_2O(g)$ ammonium nitrate  $\varepsilon_r$  dinitrogen monoxide CH<sub>2</sub>CO<sub>2</sub>H(aq) + 2 O<sub>2</sub>(g) → 2 CO<sub>2</sub>(g) + 2 H<sub>2</sub>O(t)
 carbon dioxide 3. 2 Pb(NO<sub>3</sub>)2(s)  $\rightarrow$  2 N2O<sub>4</sub>(g) + 2 PbO(s) + O<sub>2</sub>(g) lead(II) nitrate & dinitrogen tetraxide & lead(II) oxide 4.  $SiO_2(s) + 4$  HF  $\rightarrow$   $SiF_4(g) + 2$  H<sub>2</sub>O(t) silicon dioxide E<sub>f</sub> silicon tetrafluoride 5. \_8\_ H<sub>2</sub>S(aq) + \_8\_ Cl<sub>2</sub>(g)  $\rightarrow$  \_\_\_\_ S<sub>8</sub>(s) + \_16\_ HCl(aq) 6.  $\underline{\hspace{0.5cm}}$   $H_2CO_3(aq) + \underline{\hspace{0.5cm}}$   $\underline{\hspace{0.5cm}}$   $\underline{\hspace{0.5cm}}$ 7.  $P_4(s) + 6 \quad Br_2(t) \rightarrow 4 \quad PBr_3(t)$ phosphorous
8.  $S_8(t) + 4 \quad Cl_2(g) \rightarrow 4 \quad S_2Cl_2(g)$ disulfur dichloride 2 BsHs(ℓ) + 12 O₂(g) → 5 B₂O₂(s) + 9 H₂O(g) pentaboron nonahydride & diboron trioxide 10. 3  $H_2SO_4(aq) + 2 Al(OH)_2(s) \rightarrow Al_2(SO_4)_3(aq) + 6 H_2O(\ell)$ aluminum hydroxide & aluminum sulfate 11. 2 Al(s) + 3 Be(t)  $\rightarrow$  2 AlBra(s) aluminum bromide 12.  $\underline{\hspace{0.5cm}}$  (NH4)2Cr2O7(s)  $\rightarrow$   $\underline{\hspace{0.5cm}}$  N2(g) +  $\underline{\hspace{0.5cm}}$  Cr2O3(s) +  $\underline{\hspace{0.5cm}}$  H2O( $\ell$ ) annonium dichromate  $\ell$ r chromium(III) oxide 13.  $S_8(s) + 24 F_2(g) \rightarrow 8 SF_6(g)$ sulfer hexafluoride 14.  $\underline{\hspace{1cm}}$  Zn(s) +  $\underline{\hspace{1cm}}$  Pb(NO<sub>3</sub>)<sub>2</sub>(aq)  $\rightarrow$   $\underline{\hspace{1cm}}$  Zn(NO<sub>3</sub>)<sub>2</sub>(aq) +  $\underline{\hspace{1cm}}$  Pb(s)

# **Balancing Chemical Equations Multiple Choice Test**

**James T. Shipman** 

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