

Chapter 10 Review: Chemical Quantities & Calculations

(aka – mole & volume calculations, % composition, empirical & molecular formulas)

For the test, be sure that you are familiar, and able to complete, the problems below as well as the problems found on the following worksheets: "Molecular Weight and Mole Calculations," "Percent Comp. & Chemical Formula Calculations" and "Chapter 10 Sample Problems." Check the class website for answers to the review worksheet.

1. The air around you is made up of 20.95% oxygen. How many oxygen molecules are in 75.0 grams of air? The density of air is 1.19 g/L, *assume STP*.
2. How much helium (*grams*) would be needed to inflate a balloon to a volume of 5.50L? *assume STP*.
3. The molecular formula of an antibacterial drug is $C_{17}H_{15}FN_2O_5$. How many fluorine atoms are in 150mg of this drug?
4. A molecule is made up of 40.0% carbon, 6.67% hydrogen, and an unknown amount of oxygen. What is the molecular formula of the molecule if its molar mass is 90.0g/mole?
5. What is the empirical formula for a compound made up of 71.72% chlorine, 16.16% oxygen and 12.12% carbon? Calculate the molecular formula for the compound if it has a molar mass of 296.7g/mol.
6. How many molecules are in 8.42 Kilograms of C_2H_4 ?

Ch 10 Chemical Quantities Practice Problems

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Staley, Michael S. Matta, Edward L.
Waterman, Prentice-Hall Staff**



Ch 10 Chemical Quantities Practice Problems:

General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

Ebook: Introductory Chemistry: An Atoms First Approach Burdge, 2016-04-16 Ebook Introductory Chemistry An Atoms First Approach **Chemistry** John S. Phillips, Cheryl Wistrom, 2000 *Addison-Wesley Chemistry*, 1990 **Addison-Wesley Chemistry** Antony C.

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The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value

Crystal Structures M Ladd, 1999-06-01 This survey of the important types of inorganic and organic crystal structures treats its subject thoroughly and in sufficient depth for undergraduate modules in chemistry courses Features of this book are the instructions for 3D stereoviewing which is central to a full appreciation of the presentation Clear directions for making your own stereo have been provided in the book which enables readers to examine the plentiful stereo of lattices and crystal structures which are illustrated The introductory chapter explains point group and space group symmetry insofar as required to understand lattices and crystal structures Crystal structures are subdivided according to the atomic force mainly responsible for cohesion in the solid state The descriptions of the structures are given in crystallographic terms including data

on the space group molecular symmetry and molecular geometry Discussions of bonding theory for each sub division of the structures enhance and strengthen the author s presentation The book stems from the author s successful lecture courses tested and refined in class teaching It draws as necessary on equilibrium thermodynamics and other chemical topics with avoidance of advanced mathematics A level being the prerequisite Examines the important types of inorganic and organic crystal structures Includes instructions for making simple stereoviewers and computer programs Draws as necessary on equilibrium thermodynamics and other chemical topics with avoidance of advanced mathematics Merrill Chemistry

Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Fundamentals of Chemistry** David E. Goldberg, 2001

Problems for Introductory University Chemistry James Newton Butler, Basil A. Dunell, Lionel G. Harrison, 1967 The Chemical Age , 1923 **Introductory Chemistry** Darrell D. Ebbing, Albert E. Richardson, 1995 **The Chemical News**

and Journal of Industrial Science , 1921 **The English Catalogue of Books ...** Sampson Low, 1893 Chemical News

and Journal of Physical Science William Crookes, James H. Gardiner, Gerald Druce, H. W. Blood-Ryan, 1912 The Chemical

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