

Chemistry 12 – Unit I – Reaction Kinetics
Kinetics Review Test

1. Consider the following reaction mechanism:

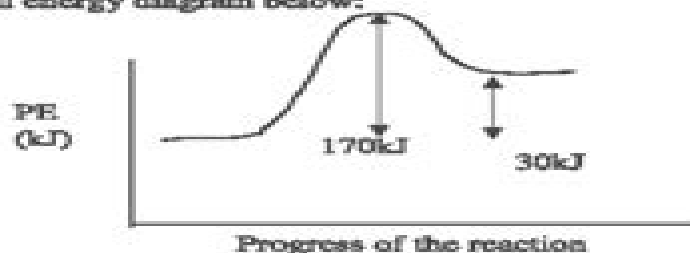


The chemical species MX is a(n) _____.

- A. catalyst B. inhibitor C. product D. reaction intermediate
2. Consider the following reaction:

$$2 \text{N}_2\text{O}_5(g) \rightarrow 4 \text{NO}_2(g) + \text{O}_2(g)$$
 At a certain temperature the rate of decomposition of $\text{N}_2\text{O}_{5(g)}$ is $2.5 \times 10^{-6} \text{ mol/s}$. The rate of formation of NO_2 is _____.
- A. $1.0 \times 10^{-5} \text{ mol/s}$ B. $1.3 \times 10^{-6} \text{ mol/s}$ C. $2.5 \times 10^{-6} \text{ mol/s}$ D. $5.0 \times 10^{-6} \text{ mol/s}$
3. Which of the following factors affect the rates of *both* homogeneous and heterogeneous reactions?
- | | |
|--------------------------------|-----------------------------------|
| I nature of the reactants | II presence of a catalyst |
| III temperature of system | IV concentration of reactants |
- A. I and IV only B. II and III only C. II, III, and IV only D. I, II, III, and IV
4. Which of the following equations represents an endothermic reaction?
- A. $\text{N}_2\text{O}_{5(g)} + 59 \text{ kJ} \rightarrow 2 \text{NO}_{2(g)}$ B. $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} + 572 \text{ kJ}$
 C. $2 \text{BrCl}_{(g)} - 29.3 \text{ kJ} \rightarrow \text{Br}_{2(g)} + \text{Cl}_{2(g)}$ D. $2 \text{H}_{2(g)} + \text{O}_{2(g)} \rightarrow 2 \text{H}_2\text{O}_{(g)} ; \Delta H = -572 \text{ kJ}$

5. Consider the potential energy diagram below:



The activation energy for the reverse reaction is _____.

- A. 30 kJ B. 140 kJ C. 170 kJ D. 200 kJ

Ch 17 Chemistry Reaction Kinetics Practice Test

John T. Moore, Richard H. Langley



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