Balancing Equations: Answers to Practice Problems

1. Balanced equations. (Coefficients equal to one (1) do not need to be shown in your answers).

(b)
$$4 \operatorname{Fe} + 3 \operatorname{O}_2 \longrightarrow 2 \operatorname{Fe}_2 \operatorname{O}_3$$

(c)
$$2 \operatorname{FeBr}_3 + 3 \operatorname{H}_2 \operatorname{SO}_4 \longrightarrow 1 \operatorname{Fe}_2(\operatorname{SO}_4)_3 + 6 \operatorname{HBr}$$

(d)
$$1 C_4H_6O_3 + 1 H_2O \longrightarrow 2 C_2H_4O_2$$

(e)
$$1 C_2H_4 + 3 O_2 \longrightarrow 2 CO_2 + 2 H_2O$$

(g)
$$1 C_7 H_{16} + 11 O_2 \longrightarrow 7 CO_2 + 8 H_2 O$$

(h)
$$4 H_2 SiCl_2 + 4 H_2 O \longrightarrow 1 H_8 Si_4 O_4 + 8 HCl$$

(i)
$$10 \, HSiCl_3 + 15 \, H_2O \longrightarrow 1 \, H_{10}Si_{10}O_{15} + 30 \, HCl$$

(j)
$$1 C_7 H_0 + 3 HNO_1 \longrightarrow 1 C_7 H_0 (NO_5)_1 + 3 H_7 O$$

(k)
$$1 C_5 H_8 O_2 + 2 NaH + 2 HCI \longrightarrow 1 C_5 H_{12} O_2 + 2 NaCl$$

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Thomas Griffiths

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