

# 3-1 Skills Practice

## Parallel Lines and Transversals

For Exercises 1–4, refer to the figure at the right.

1. Name all planes that are parallel to plane  $DEH$ . plane  $FBG$
2. Name all segments that are parallel to  $\overline{AB}$ .  $\overline{CD}$ ,  $\overline{EF}$ ,  $\overline{GH}$
3. Name all segments that intersect  $\overline{GH}$ .  $\overline{CG}$ ,  $\overline{DH}$ ,  $\overline{EH}$ ,  $\overline{FG}$
4. Name all segments that are skew to  $\overline{CD}$ .  $\overline{AE}$ ,  $\overline{BF}$ ,  $\overline{EH}$ ,  $\overline{FG}$



Lesson 3-1

Identify the sets of lines to which each given line is a transversal.

5.  $r$   $s$  and  $t$ ,  $s$  and  $w$ ,  $s$  and  $z$ ,  $t$  and  $w$ ,  $t$  and  $z$ ,  $w$  and  $z$
6.  $s$   $r$  and  $t$ ,  $r$  and  $w$ ,  $r$  and  $z$ ,  $t$  and  $w$ ,  $t$  and  $z$ ,  $w$  and  $z$
7.  $w$   $r$  and  $s$ ,  $r$  and  $t$ ,  $r$  and  $z$ ,  $s$  and  $t$ ,  $s$  and  $z$ ,  $t$  and  $z$



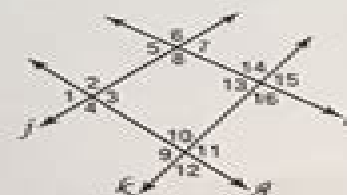
Identify each pair of angles as *alternate interior*, *alternate exterior*, *corresponding*, or *consecutive interior* angles.

- |  |  |
|--|--|
| 8. $\angle 2$ and $\angle 8$<br>alternate exterior   | 9. $\angle 3$ and $\angle 6$<br>consecutive interior |
| 10. $\angle 1$ and $\angle 9$<br>corresponding       | 11. $\angle 3$ and $\angle 9$<br>alternate interior  |
| 12. $\angle 6$ and $\angle 12$<br>alternate exterior | 13. $\angle 7$ and $\angle 11$<br>corresponding      |



Name the transversal that forms each pair of angles. Then identify the special name for the angle pair.

- |  |   |
|--|---|
| 14. $\angle 4$ and $\angle 10$<br>$g$ ; alternate interior | 15. $\angle 2$ and $\angle 12$<br>$g$ ; alternate exterior    |
| 16. $\angle 7$ and $\angle 3$<br>$j$ ; corresponding       | 17. $\angle 13$ and $\angle 10$<br>$k$ ; consecutive interior |
| 18. $\angle 8$ and $\angle 14$<br>$h$ ; alternate interior | 19. $\angle 6$ and $\angle 14$<br>$h$ ; corresponding         |



# Ch 10 Test Mcdougal Littel Geometry Answers

**Albert A Gayle**



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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the

titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1  
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$  1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak  $pK_a$  of an unknown acid or  $pK_b$  of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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web conclusion in this lab we used titration to explore the concepts of stoichiometry and equivalence points we found the concentration of an unknown substance by mixing 2m hcl with the naoh of unknown concentration in order to experimentally ascertain the concentration of the naoh

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  $\text{NaOH(aq)} + \text{C}_6\text{H}_4\text{COOH(aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa(aq)} + \text{H}_2\text{O(l)}$  the point of equivalence is the point of titration when the number of moles of naoh na added is equal to the number of moles of khp nb in the

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